

Chapter 13 - Practical Work

(a) (i) Chemical test for ammonia:

If a rod dipped in concentrated hydrochloric acid is brought near ammonia gas, dense white fumes of ammonium chloride (NH_4Cl) are formed.

(ii) Chemical test for Sulphur dioxide:

It decolorizes pink coloured potassium permanganate solution.



(iii) Chemical test for HCl:

When HCl gas is passed through AgNO_3 solution, white precipitates of AgCl are formed which gets dissolved in excess of NH_4OH .



(iv) Chemical test for Chlorine:

It turns moist starch iodide paper (KI + starch solution) blue black.

(v) Chemical test for Carbon dioxide:

When this gas is passed through lime water, it turns milky due to the formation of white precipitates of CaCO_3 and on passing excess of carbon dioxide gas, this milkiness disappears.

(vi) Chemical test for oxygen:

This gas is absorbed in colourless alkaline solution of pyrogallol and turns it dark brown.

(vii) Chemical test for hydrogen:

It burns with a pop sound when a burning taper is brought near it.

(b) Ammonia is a basic gas and its basic nature is suspected through litmus paper test because it changes the colour of red litmus paper to blue.

(c) Chlorine, carbon dioxide, hydrogen chloride, hydrogen sulphide and sulphur dioxide are acidic gases since they convert blue litmus to red.

(d) A is chlorine and B is Sulphur dioxide.

(e) Water vapour.

Chemistry Chapter 13 Solutions Manual

Susan McMurry



Chemistry Chapter 13 Solutions Manual:

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Table of Contents Chemistry Chapter 13 Solutions Manual

1. Understanding the eBook Chemistry Chapter 13 Solutions Manual
 - The Rise of Digital Reading Chemistry Chapter 13 Solutions Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 13 Solutions Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Chapter 13 Solutions Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 13 Solutions Manual
 - Personalized Recommendations

- Chemistry Chapter 13 Solutions Manual User Reviews and Ratings
- Chemistry Chapter 13 Solutions Manual and Bestseller Lists
- 5. Accessing Chemistry Chapter 13 Solutions Manual Free and Paid eBooks
 - Chemistry Chapter 13 Solutions Manual Public Domain eBooks
 - Chemistry Chapter 13 Solutions Manual eBook Subscription Services
 - Chemistry Chapter 13 Solutions Manual Budget-Friendly Options
- 6. Navigating Chemistry Chapter 13 Solutions Manual eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Chapter 13 Solutions Manual Compatibility with Devices
 - Chemistry Chapter 13 Solutions Manual Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Chapter 13 Solutions Manual
 - Highlighting and Note-Taking Chemistry Chapter 13 Solutions Manual
 - Interactive Elements Chemistry Chapter 13 Solutions Manual
- 8. Staying Engaged with Chemistry Chapter 13 Solutions Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 13 Solutions Manual
- 9. Balancing eBooks and Physical Books Chemistry Chapter 13 Solutions Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 13 Solutions Manual
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Chapter 13 Solutions Manual
 - Setting Reading Goals Chemistry Chapter 13 Solutions Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Chapter 13 Solutions Manual
 - Fact-Checking eBook Content of Chemistry Chapter 13 Solutions Manual

- Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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