

Student: Noor Alsamara

## Questions

Separate ions in unknown solution: You are given an unknown solution, which contains both aqueous barium ion,  $\text{Ba}^{2+}(\text{aq})$ , and aqueous lead(II) ion,  $\text{Pb}^{2+}(\text{aq})$ . Which of your six solutions would allow separation of the ions?

Select all that apply.

Sodium chloride,  $\text{NaCl}(\text{aq})$

Sodium hydroxide,  $\text{NaOH}(\text{aq})$

## Collected Lab Data

### Describe solutions 1 through 6

Solution	Color	Transparency
Solution 1	pink	clear
Solution 2	colorless	clear
Solution 3	colorless	clear
Solution 4	colorless	clear
Solution 5	colorless	clear
Solution 6	colorless	clear

### Observe reactions with solutions 1 through 6

Mixture	Observations (Select all that apply)	Precipitate color
Solutions 1 + 2	Precipitate formed	Dark pink
Solutions 1 + 3	Precipitate formed	Pale pink
Solutions 1 + 4	No reaction occurred	No precipitate formed
Solutions 1 + 5	No reaction occurred	No precipitate formed
Solutions 1 + 6	No reaction occurred	No precipitate formed
Solutions 2 + 3	No reaction occurred	No precipitate formed
Solutions 2 + 4	Gas produced Test tube became hotter	No precipitate formed
Solutions 2 + 5	No reaction occurred	No precipitate formed

# Chemistry 12 Lab Qualitative Analysis 16e Answers

**Wenbin Ji**



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