

## Chapter Test B

### Chapter: Chemical Equations and Reactions

Write the correct number in the space provided with the letter of the unit, if given. If none is given, please use the appropriate SI unit. Write answers with units.

- The molar volume of an ideal gas at STP is  $22.71 \text{ L mol}^{-1}$ .
  - gas
  - grams
  - moles
  - liters
- The volume of a gas at STP is  $1.00 \text{ L}$ . How many moles of gas are there?
  - moles
  - liters
  - grams
  - moles of particles
- An unknown gas has a density of  $1.96 \text{ g L}^{-1}$  at STP. What is the molar mass?
  - grams
  - liters
  - moles
  - moles
- The molar mass of a gas is  $44.01 \text{ g mol}^{-1}$ . How many moles of gas are there if the mass is  $1.00 \text{ g}$ ?
  - grams
  - liters
  - moles
  - moles of particles
- The molar mass of a gas is  $44.01 \text{ g mol}^{-1}$ . How many moles of gas are there if the mass is  $1.00 \text{ g}$ ?
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  - liters
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  - moles of particles
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  - moles of particles

# Chapter Test B Chemical Equations And Reactions Answers

**Martha Mackin**



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