

Chapter 7 Basics of Chemistry

1. The science that deals with the composition, structure, and properties of matter and how matter changes under different conditions
2. The study of substances that contain the element carbon
3. The branch of chemistry dealing with compounds that do not contain the element carbon, but may contain the element hydrogen
4. Any substance that occupies space and has mass
5. The simplest form of chemical matter; can't be broken down without a loss of identity
6. The particles from which all matter is composed
7. Formed by joining two or more atoms chemically
8. Contain two or more atoms of the same element that are united chemically
9. Chemical combinations of two or more atoms of different elements that are united chemically
10. Have definite size and a definite shape
11. Have a definite size but a definite shape
12. Do not have a definite size or a definite shape
13. Characteristics that can be determined without a chemical reaction and that do not cause a chemical change in the identity of the substances
14. Characteristics that can be determined only with a chemical reaction and that cause a chemical change in the identity of the substance
15. Change in the form or physical properties of a substance without the formation of a new substance
16. Change in the chemical composition of a substance, in which a new substance are formed having properties different from the original
17. Combination of two or more atoms of different elements united chemically; fixed chemical composition, definite proportions, distinct properties
18. Combination of two or more substances that are united physically, in any proportions with combined properties
19. (H) colorless, odorless, tasteless gas and the lightest element known
20. (O) The most abundant element found on earth, is a colorless, odorless, tasteless gas

- A. Hydrogen
- B. Organic chemistry
- C. Chemical properties
- D. Elemental molecules
- E. Liquids
- F. Physical mixtures
- G. Nitrogen
- H. Physical change
- I. Compound molecules
- J. Oxygen
- K. Air
- L. Chemical compounds
- M. Physical properties
- N. Hydrogen peroxide
- O. Chemical change
- P. Gases
- Q. Molecule
- R. Chemistry
- S. Solids
- T. Elements

Chapter 7 Chemistry Test

Thomas Evangelist



Chapter 7 Chemistry Test:

McGraw-Hill's SAT Subject Test Chemistry, 3rd Edition Thomas Evangelist, 2011-02-03 Expert guidance on the Chemistry exam Many colleges and universities require you to take one or more SAT II Subject Tests to demonstrate your mastery of specific high school subjects McGraw Hill's SAT Subject Test Chemistry is written by experts in the field and gives you the guidance you need to perform at your best This book includes 4 full length sample tests updated for the latest test format 40 top tips to remember on test day Glossary of tested chemistry terms and formulas Tips and strategies from one of the most popular teachers at the renowned Brooklyn Technical High School Diagnostic test to pinpoint strengths and weaknesses Step by step review of all topics covered on the exam In depth coverage of the lab experiment questions that are a major test feature Charts tables and illustrations to simplify and reinforce learning Practice tests just like the real SAT Subject Test in Chemistry Test taking tips and strategies

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In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Chapter 7 Chemistry Test**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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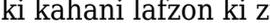
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