

Chapter 11 Nuclear Chemistry

Review

Introduction

Chemical Energy is energy stored in the bonds of atoms and molecules. **Nuclear Energy** is stored through the nucleus of atoms changing their their different nuclei. Nuclear energy is produced by **radioactivity**.

The Nucleus

Composition of Nucleus



Atomic Number

Nucleus is composed of two main particles, protons and neutrons.

The number of protons is the atomic number.

The number of protons and neutrons together is the mass number.

Neutrons that come from the nucleus are called **neutrons**. The number of neutrons in any atom is:

Mass number - **atomic number** = no. of neutrons

Example: ${}^{23}_{11}\text{Na}$ has 12 neutrons.

Table 11.1: Properties of alpha, beta and gamma rays			
	Alpha rays	Beta rays	Gamma rays
Symbol	α	β	γ
Charge	2+	1-	0
Mass	4 u	$\frac{1}{1836}$ u	0
Penetrating power	Low	High	Very high
Ionising power	High	Low	Very low

Types of Radioactive Decay

1. **Alpha decay** is the loss of an α -particle.



2. **Beta decay** is the loss of a β -particle (a high-speed electron) from the nucleus: ${}^A_ZX \rightarrow {}^A_{Z+1}Y$



3. **Gamma radiation** is the loss of a γ -ray, which is a high-energy electromagnetic wave. Gamma rays' emission does not change the mass or the atomic number of the nucleus: ${}^A_ZX \rightarrow {}^A_ZX + \gamma$

Chapter 21 Nuclear Chemistry Mixed Review Answers

Michael Hoelscher



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Table of Contents Chapter 21 Nuclear Chemistry Mixed Review Answers

1. Understanding the eBook Chapter 21 Nuclear Chemistry Mixed Review Answers
 - The Rise of Digital Reading Chapter 21 Nuclear Chemistry Mixed Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 21 Nuclear Chemistry Mixed Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 21 Nuclear Chemistry Mixed Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 21 Nuclear Chemistry Mixed Review Answers
 - Personalized Recommendations
 - Chapter 21 Nuclear Chemistry Mixed Review Answers User Reviews and Ratings

- Chapter 21 Nuclear Chemistry Mixed Review Answers and Bestseller Lists
- 5. Accessing Chapter 21 Nuclear Chemistry Mixed Review Answers Free and Paid eBooks
 - Chapter 21 Nuclear Chemistry Mixed Review Answers Public Domain eBooks
 - Chapter 21 Nuclear Chemistry Mixed Review Answers eBook Subscription Services
 - Chapter 21 Nuclear Chemistry Mixed Review Answers Budget-Friendly Options
- 6. Navigating Chapter 21 Nuclear Chemistry Mixed Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 21 Nuclear Chemistry Mixed Review Answers Compatibility with Devices
 - Chapter 21 Nuclear Chemistry Mixed Review Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 21 Nuclear Chemistry Mixed Review Answers
 - Highlighting and Note-Taking Chapter 21 Nuclear Chemistry Mixed Review Answers
 - Interactive Elements Chapter 21 Nuclear Chemistry Mixed Review Answers
- 8. Staying Engaged with Chapter 21 Nuclear Chemistry Mixed Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 21 Nuclear Chemistry Mixed Review Answers
- 9. Balancing eBooks and Physical Books Chapter 21 Nuclear Chemistry Mixed Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 21 Nuclear Chemistry Mixed Review Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 21 Nuclear Chemistry Mixed Review Answers
 - Setting Reading Goals Chapter 21 Nuclear Chemistry Mixed Review Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 21 Nuclear Chemistry Mixed Review Answers
 - Fact-Checking eBook Content of Chapter 21 Nuclear Chemistry Mixed Review Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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