

Study Card: Strong vs. Weak Acids and Bases

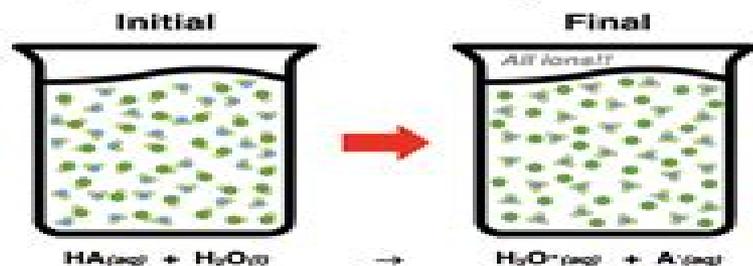
Acids and bases dissociate into ions in aqueous solutions. Whether an acid or a base is classified as strong or weak depends on the **degree of dissociation**.

Strong Acids: completely dissociate into ions when dissolved in water.



Goes to completion.

Equilibrium lies far to the right. $K \gg \gg 1$.

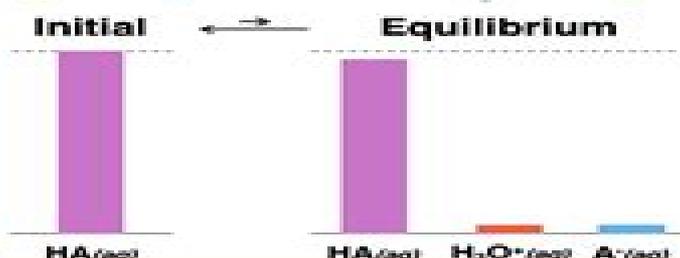
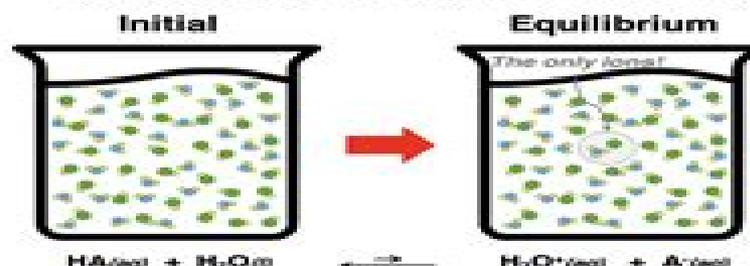


Weak Acids: partially dissociate into ions when dissolved in water.



Reversible; with a strong reverse reaction.

Equilibrium lies far to the left. $K \ll \ll 1$.



Examples

Strong Acids	Strong Bases
Hydrochloric acid (HCl)	Lithium hydroxide (LiOH)
Sulfuric acid (H ₂ SO ₄)	Sodium hydroxide (NaOH)
Nitric acid (HNO ₃)	Potassium hydroxide (KOH)
Hydrobromic acid (HBr)	Rubidium hydroxide (RbOH)
Hydroiodic acid (HI)	Cesium hydroxide (CsOH)
Perchloric acid (HClO ₄)	Calcium hydroxide (Ca(OH) ₂)
Chloric acid (HClO ₃)	Strontium hydroxide (Sr(OH) ₂)
	Barium hydroxide (Ba(OH) ₂)

Need help with calculations of the pH of strong acids or bases? You'll find help and examples on our website.

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