

Calculating Average Atomic Mass Worksheet:

- 1) Three isotopes of Silicon occur in nature:

Isotopes of Silicon:	Percent Abundance:	Atomic Mass:
Silicon-28	92.23%	27.97693 amu
Silicon-29	4.68%	28.97649 amu
Silicon-30	3.09%	29.97377 amu

Calculate the average atomic mass for the three isotopes of Silicon.

$$(27.97693 \times 0.9223) + (28.97649 \times 0.0468) + (29.97377 \times 0.0309) = 28.0854 \text{ amu}$$

Si-28 Si-29 Si-30

- 2) Two isotopes of Rubidium occur naturally:

Isotopes of Rubidium:	Percent Abundance:	Atomic Mass:
Rubidium-85	72.15%	84.9118 amu
Rubidium-87	27.85%	86.9092 amu

Calculate the average atomic mass for the two isotopes of Rubidium.

$$(84.9118 \times 0.7215) + (86.9092 \times 0.2785) = 85.468$$

Rubidium-85 Rubidium-87

- 3) Only two isotopes of Copper naturally occur:

Isotopes of Copper:	Percent Abundance:	Atomic Mass:
Copper-63	69.17%	62.9396 amu
Copper-65	30.83%	64.9278 amu

Calculate the average atomic mass for the two isotopes of Copper.

$$(62.9396 \times 0.6917) + (64.9278 \times 0.3083) = 63.552$$

Copper-63 Copper-65

Calculating Average Atomic Mass Worksheet Answers

Tao Wei



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