

## Balancing Chemical Equations

Once we have written a chemical equation, we must make sure to balance it! Balancing a chemical equation means to make sure that there are the same number of each type of atom on both sides of the chemical equation. So, if there are five carbons on the left side (reactants), there needs to be five carbons on the right side (products).

So, how do we balance an equation? Follow the following steps...

- Change **only** stoichiometric coefficients – those big numbers in front of the compounds (not subscripts in the formulas – that would change the identity of the compound!)
- Start with most complex
- If found in more than one place on each side, do it last
- Watch out for odd numbers (make them into even ones if needed)
- Check your work!

Now, try to balance the following equations. In addition, name the **bolded** compounds for some nomenclature practice!

- $\underline{\quad} \text{NH}_4\text{NO}_3(\text{s}) \rightarrow \underline{\quad} \text{N}_2\text{O}(\text{g}) + \underline{2} \text{H}_2\text{O}(\text{g})$   
*ammonium nitrate & dinitrogen monoxide*
- $\underline{\quad} \text{CH}_3\text{CO}_2\text{H}(\text{aq}) + \underline{2} \text{O}_2(\text{g}) \rightarrow \underline{2} \text{CO}_2(\text{g}) + \underline{2} \text{H}_2\text{O}(\text{l})$   
*oxygen & carbon dioxide*
- $\underline{2} \text{Pb}(\text{NO}_3)_2(\text{s}) \rightarrow \underline{2} \text{N}_2\text{O}_4(\text{g}) + \underline{2} \text{PbO}(\text{s}) + \underline{\quad} \text{O}_2(\text{g})$   
*lead(II) nitrate & dinitrogen tetroxide & lead(II) oxide*
- $\underline{\quad} \text{SiO}_2(\text{s}) + \underline{4} \text{HF} \rightarrow \underline{\quad} \text{SiF}_4(\text{g}) + \underline{2} \text{H}_2\text{O}(\text{l})$   
*silicon dioxide & silicon tetrafluoride*
- $\underline{8} \text{H}_2\text{S}(\text{aq}) + \underline{8} \text{Cl}_2(\text{g}) \rightarrow \underline{\quad} \text{S}_8(\text{s}) + \underline{16} \text{HCl}(\text{aq})$   
*chlorine*
- $\underline{\quad} \text{H}_2\text{CO}_3(\text{aq}) + \underline{\quad} \text{Ba}(\text{OH})_2 \rightarrow \underline{\quad} \text{BaCO}_3(\text{s}) + \underline{2} \text{H}_2\text{O}(\text{l})$   
*barium hydroxide & barium carbonate*
- $\underline{\quad} \text{P}_4(\text{s}) + \underline{6} \text{Br}_2(\text{l}) \rightarrow \underline{4} \text{PBr}_3(\text{l})$   
*phosphorous & phosphorous tribromide*
- $\underline{\quad} \text{S}_8(\text{l}) + \underline{4} \text{Cl}_2(\text{g}) \rightarrow \underline{4} \text{S}_2\text{Cl}_2(\text{g})$   
*disulfur dichloride*
- $\underline{2} \text{B}_5\text{H}_9(\text{l}) + \underline{12} \text{O}_2(\text{g}) \rightarrow \underline{5} \text{B}_2\text{O}_3(\text{s}) + \underline{9} \text{H}_2\text{O}(\text{g})$   
*pentaboron nonahydride & diboron trioxide*
- $\underline{3} \text{H}_2\text{SO}_4(\text{aq}) + \underline{2} \text{Al}(\text{OH})_3(\text{s}) \rightarrow \underline{\quad} \text{Al}_2(\text{SO}_4)_3(\text{aq}) + \underline{6} \text{H}_2\text{O}(\text{l})$   
*aluminum hydroxide & aluminum sulfate*
- $\underline{2} \text{Al}(\text{s}) + \underline{3} \text{Br}_2(\text{l}) \rightarrow \underline{2} \text{AlBr}_3(\text{s})$   
*aluminum bromide*
- $\underline{\quad} (\text{NH}_4)_2\text{Cr}_2\text{O}_7(\text{s}) \rightarrow \underline{\quad} \text{N}_2(\text{g}) + \underline{\quad} \text{Cr}_2\text{O}_3(\text{s}) + \underline{4} \text{H}_2\text{O}(\text{l})$   
*ammonium dichromate & chromium(III) oxide*
- $\underline{\quad} \text{S}_8(\text{s}) + \underline{24} \text{F}_2(\text{g}) \rightarrow \underline{8} \text{SF}_6(\text{g})$   
*sulfur hexafluoride*
- $\underline{\quad} \text{Zn}(\text{s}) + \underline{\quad} \text{Pb}(\text{NO}_3)_2(\text{aq}) \rightarrow \underline{\quad} \text{Zn}(\text{NO}_3)_2(\text{aq}) + \underline{\quad} \text{Pb}(\text{s})$   
*zinc nitrate*

# Balancing Chemical Equations Multiple Choice Test

**Arturo Cuomo**



**Balancing Chemical Equations Multiple Choice Test:**

## Decoding **Balancing Chemical Equations Multiple Choice Test**: Revealing the Captivating Potential of Verbal Expression

In an era characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its power to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Balancing Chemical Equations Multiple Choice Test**," a mesmerizing literary creation penned by a celebrated wordsmith, readers attempt an enlightening odyssey, unraveling the intricate significance of language and its enduring affect our lives. In this appraisal, we shall explore the book is central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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