

PHYSICAL PHARMACY

PHYSICAL CHEMISTRY – branch of chemistry concerned with physical properties and structure of matter and with laws and theories of physical and chemical changes

PHYSICAL PHARMACY – application of physical chemistry in pharmacy; study of physicochemical properties of substances used in drug formulation

I. FORCES OF ATTRACTION

A. INTRAMOLECULAR – intra (within) the molecule

- **Ionic Bond** – transfer of electrons; metal (electron donor) + non-metal (electron acceptor); (ex. Na: $p^+ = 11$; $e^- = 11 \rightarrow p^+ = 11$; $e^- = 10 \rightarrow$ charge = +1)

○ PERIODIC TRENDS REPRESENTING BONDS

- **Ionization Energy** – energy required remove electron (low M; high NM)
- **Electronegativity** – ability to attract electrons (low M; high NM)

- **Covalent Bonds** – sharing of electrons; non-metal + non-metal; polarity – difference in electronegativity (symmetrical = non-polar)

- **Polar** – unequal sharing of electrons; dipole moment – (ex. $H^{\delta+}-Br^{\delta-}$)
- **Non Polar** – equal sharing of electrons (ex. $Br-Br$)

B. INTERMOLECULAR – inter (between) the molecules

- **Van der Waals**

- **Dipole-Dipole** $\rightarrow P + P$; Keesom Forces; orientation effect
- **Dipole-Induced Dipole** $\rightarrow P + NP$; Debye Forces; inductive effect
- **Induced Dipole-Induced Dipole** $\rightarrow NP + NP$; London Forces; dispersive effect

- **Ion-Induced Dipole** $\rightarrow ion + NP$

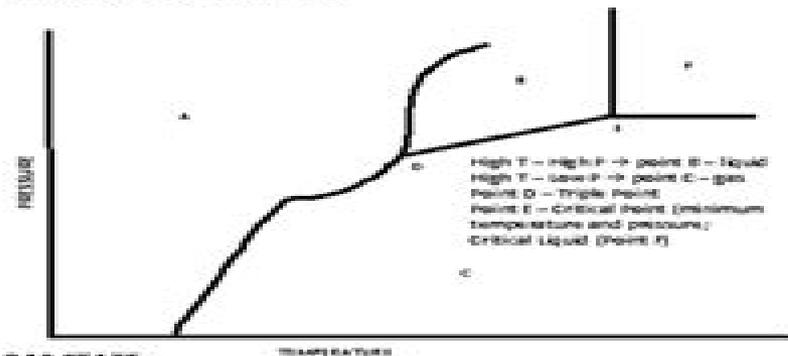
$I_2 \rightarrow NP$ (solubility = 1:2960)

$I_2 + NaI/KI$ (solubilizing agent) $\rightarrow I_2 + I_3^-$ = increase solubility

- **Hydrogen Bond** – hydrogen atom + electronegative atom (N, O, F)



PHASE DIAGRAM



GAS STATE:

- **Kinetic Molecular Theory**

- Gases are composed of small particles called atoms; total volume is negligible in relation to space where they are confined
- Particles do not attract one another but move with complete independence
- Particles exhibit continuous random motion owing to their kinetic energy
- Gases exhibit perfect elasticity (no energy exchange)

II. GAS LAWS

- at standard temperature (Kelvin = $^{\circ}C + 273$) and pressure (1 atm = 760 mmHg or torr; 760 mmHg; 1.01325×10^5 dynes/cm²; 1.01325×10^5 Pa or N/m²; 0.1325 bar)

- **Boyle's Law** – constant temperature; increase P = decrease V; inversely related

$$P \propto \frac{1}{V}$$

- * density = increase altitude; decrease volume; $P_1V_1 = P_2V_2$

- **Charles' Law** – constant pressure; increase T = increase V

$$T \propto V; \frac{T_1}{T_2} = \frac{V_1}{V_2} \text{ or } \frac{T_1}{V_1} = \frac{T_2}{V_2}$$

- * H_2O – 1.0g/ml $\rightarrow 4^{\circ}C$

- **Gay Lussac's Law** \rightarrow constant volume

$$P \propto T; \frac{P_1}{P_2} = \frac{T_1}{T_2} \text{ or } \frac{P_1}{T_1} = \frac{P_2}{T_2}$$

- * $121^{\circ}C$ – 15-20psi

Physical Pharmacy Lecture Notes

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Table of Contents Physical Pharmacy Lecture Notes

1. Understanding the eBook Physical Pharmacy Lecture Notes
 - The Rise of Digital Reading Physical Pharmacy Lecture Notes
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Pharmacy Lecture Notes
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Pharmacy Lecture Notes
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Pharmacy Lecture Notes

- Personalized Recommendations
 - Physical Pharmacy Lecture Notes User Reviews and Ratings
 - Physical Pharmacy Lecture Notes and Bestseller Lists
5. Accessing Physical Pharmacy Lecture Notes Free and Paid eBooks
 - Physical Pharmacy Lecture Notes Public Domain eBooks
 - Physical Pharmacy Lecture Notes eBook Subscription Services
 - Physical Pharmacy Lecture Notes Budget-Friendly Options
 6. Navigating Physical Pharmacy Lecture Notes eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Pharmacy Lecture Notes Compatibility with Devices
 - Physical Pharmacy Lecture Notes Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Pharmacy Lecture Notes
 - Highlighting and Note-Taking Physical Pharmacy Lecture Notes
 - Interactive Elements Physical Pharmacy Lecture Notes
 8. Staying Engaged with Physical Pharmacy Lecture Notes
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Pharmacy Lecture Notes
 9. Balancing eBooks and Physical Books Physical Pharmacy Lecture Notes
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Pharmacy Lecture Notes
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Physical Pharmacy Lecture Notes
 - Setting Reading Goals Physical Pharmacy Lecture Notes
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Physical Pharmacy Lecture Notes

- Fact-Checking eBook Content of Physical Pharmacy Lecture Notes
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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