



Name:

Date:

Student Exploration: Phase Changes (4b)

Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

Vocabulary: altitude, boil, boiling point, freeze, freezing point, gas, liquid, melt, melting point, phase, solid

Activity B:

Temperature and molecular motion

Get the Gizmo ready:

- Click **Reset**, and select the **Micro view**.
- Set **Ice volume** to 0 cm³.
- Set **Add/remove heat energy** to 0 J/s.



Question: Why do phase changes occur?

1. **Compare:** Set the **Water temperature** to 0 °C and click **Play**. Observe the water molecules. Click **Reset**, set the **Water temperature** to 100 °C, and click **Play** again.

What do you notice?

There is no difference.

2. **Observe:** Click **Reset**. The **mean molecular speed** of the water molecules is displayed below the container. Set the **Water temperature** to 0 °C and **Add/remove heat energy** to 400 J/s. Click **Play**.

A. How does the mean speed of the water molecules change as they are heated?

The water gets hotter (has more heat energy).

B. Does the mean molecular speed change as much as the temperature as the water heats up? Explain.

Yes, the more the temperature increases the more the molecules of the water speed up (until they hit the boiling point, where they cannot get hotter any longer.)

3. **Explain:** How is temperature related to the motions of molecules?

The higher the temperature, the faster the molecules move.

4. **Observe:** Click **Reset**. Set the **Water temperature** to 20 °C and the **Ice volume** to 50 cm³. Set **Add/remove heat energy** to 0 J/s. Click **Play**. How do the molecules in the liquid interact with the molecules in the solid?

Phase Changes Answers Gizmo

Judd E. Hollander



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Table of Contents Phase Changes Answers Gizmo

1. Understanding the eBook Phase Changes Answers Gizmo
 - The Rise of Digital Reading Phase Changes Answers Gizmo
 - Advantages of eBooks Over Traditional Books
2. Identifying Phase Changes Answers Gizmo
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Phase Changes Answers Gizmo
 - User-Friendly Interface
4. Exploring eBook Recommendations from Phase Changes Answers Gizmo
 - Personalized Recommendations
 - Phase Changes Answers Gizmo User Reviews and Ratings
 - Phase Changes Answers Gizmo and Bestseller Lists
5. Accessing Phase Changes Answers Gizmo Free and Paid eBooks
 - Phase Changes Answers Gizmo Public Domain eBooks
 - Phase Changes Answers Gizmo eBook Subscription Services
 - Phase Changes Answers Gizmo Budget-Friendly Options

6. Navigating Phase Changes Answers Gizmo eBook Formats
 - ePub, PDF, MOBI, and More
 - Phase Changes Answers Gizmo Compatibility with Devices
 - Phase Changes Answers Gizmo Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Phase Changes Answers Gizmo
 - Highlighting and Note-Taking Phase Changes Answers Gizmo
 - Interactive Elements Phase Changes Answers Gizmo
8. Staying Engaged with Phase Changes Answers Gizmo
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Phase Changes Answers Gizmo
9. Balancing eBooks and Physical Books Phase Changes Answers Gizmo
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Phase Changes Answers Gizmo
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Phase Changes Answers Gizmo
 - Setting Reading Goals Phase Changes Answers Gizmo
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Phase Changes Answers Gizmo
 - Fact-Checking eBook Content of Phase Changes Answers Gizmo
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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