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Covalent Bonding Practice Worksheet

Explanation: Covalent compounds are compounds between nonmetal atoms. The atoms "start" with a number of electrons equal to their group number in the periodic table. Then they share electrons so that all the atoms have filled shells with a "magic number" of electrons. The magic number is 2 for hydrogen and 8 for C, N, O, F, Si, P, S, Cl, Se, Br and I. Remember that compounds made with a metal atom and a nonmetal atom make an ionic compound, not a covalent one, and compounds of metals are also different, these have metallic bonds and do not require a strict ratio.

Instructions: Fill in the blanks in the table. The first row is done for you; study the first row carefully before you start, especially the difference between different boxes (remember there are many ways to draw the same compound). Some rows have hints. If a box has an X you don't have to fill in that box.

Helpful Tips: 1) All atoms start with a number of outer shell electrons = their group number. You cannot change the element's group number. 2) If you have to choose between two shapes, the symmetrical, balanced one is usually best. 3) Try to figure the structures out yourself, don't cheat and use the Internet. ☺

Compound Formula	Full electronic structure	Lewis structure	Dot-cross diagram	Structural formula	3-d formula
H ₂ O		H:O:H ••	H:O:H ••	H-O-H	
NH ₃		H:N:H •• H	H:N:H •• H	H-N-H H	
CH ₃ Br		H:C:H •• Br	H:C:H •• Br	H-C-H Br	

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(12)

(12)

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Academy of Engineering, National
Science Resources Center of the
National Academy of
Sciences, Institute of Medicine**

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Table of Contents Core Teaching Resources Covalent Bonding Answer Key

1. Understanding the eBook Core Teaching Resources Covalent Bonding Answer Key
 - The Rise of Digital Reading Core Teaching Resources Covalent Bonding Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Core Teaching Resources Covalent Bonding Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a Core Teaching Resources Covalent Bonding Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Core Teaching Resources Covalent Bonding Answer Key
 - Personalized Recommendations
 - Core Teaching Resources Covalent Bonding Answer Key User Reviews and Ratings

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- Core Teaching Resources Covalent Bonding Answer Key and Bestseller Lists
- 5. Accessing Core Teaching Resources Covalent Bonding Answer Key Free and Paid eBooks
 - Core Teaching Resources Covalent Bonding Answer Key Public Domain eBooks
 - Core Teaching Resources Covalent Bonding Answer Key eBook Subscription Services
 - Core Teaching Resources Covalent Bonding Answer Key Budget-Friendly Options
- 6. Navigating Core Teaching Resources Covalent Bonding Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Core Teaching Resources Covalent Bonding Answer Key Compatibility with Devices
 - Core Teaching Resources Covalent Bonding Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Core Teaching Resources Covalent Bonding Answer Key
 - Highlighting and Note-Taking Core Teaching Resources Covalent Bonding Answer Key
 - Interactive Elements Core Teaching Resources Covalent Bonding Answer Key
- 8. Staying Engaged with Core Teaching Resources Covalent Bonding Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Core Teaching Resources Covalent Bonding Answer Key
- 9. Balancing eBooks and Physical Books Core Teaching Resources Covalent Bonding Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Core Teaching Resources Covalent Bonding Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Core Teaching Resources Covalent Bonding Answer Key
 - Setting Reading Goals Core Teaching Resources Covalent Bonding Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Core Teaching Resources Covalent Bonding Answer Key
 - Fact-Checking eBook Content of Core Teaching Resources Covalent Bonding Answer Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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