

Molality is number of moles of solute present in 1 kg of solvent.

e.g. if 80 g NaOH is present in 1000 g water then the molality will be calculated number of

$$\text{moles of NaOH} = \frac{\text{Amount}}{\text{Molar Mass}}$$

$$= \frac{80}{40} = 2$$

Here 2 moles of NaOH are in 1000 g (or 1 kg)

Solvent so the molality is 2 molal

$$\text{Formulae of molality} = \frac{W_{\text{Solute}}}{M_{\text{Solute}} \times W_{\text{solvent}} (\text{kg})}$$

$$= \frac{W_{\text{Solute}} \times 1000}{M_{\text{Solute}} \times W_{\text{solvent}} (g)}$$

Mol fraction: Fraction of moles of a particular component in mixture.

e.g. In mixture of He(2 mol) and Ne(3 mol) the mol fraction of He =  $\frac{2}{2+3} = \frac{2}{5} = 0.4$

the mol fraction of Ne =  $\frac{3}{2+3} = \frac{3}{5} = 0.6$

Formula for mol fraction of component A

$$\left[ X_A = \frac{n_A}{\sum_{i=1}^n n_i} = \frac{\text{number of moles of A}}{\text{Total number of moles of all species}} \right]$$

# Mole Fraction Problems And Solutions

**Oswaal Editorial Board**



## **Mole Fraction Problems And Solutions:**

## The Enigmatic Realm of **Mole Fraction Problems And Solutions**: Unleashing the Language is Inner Magic

In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing short of extraordinary. Within the captivating pages of **Mole Fraction Problems And Solutions** a literary masterpiece penned by way of a renowned author, readers set about a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book's core themes, assess its distinct writing style, and delve into its lasting impact on the hearts and minds of people who partake in its reading experience.

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