

1.1| MATTER

Anything which occupies space, has mass and can be felt by our one or more five senses is known as matter. E.g. air, water, iron etc.

CLASSIFICATION OF MATTER

Matter can be classified on the basis of their physical and chemical properties as:

BASED ON PHYSICAL PROPERTIES MATTER

- 1) Solid
- 2) Liquid
- 3) Gases

BASED ON CHEMICAL PROPERTIES

- 1) Elements
- 2) Compounds
- 3) Mixture

Note: In this chapter we shall study about matter based on their physical properties.

1.2| PHYSICAL NATURE OF MATTER

1. MATTER IS MADE UP OF VERY-VERY SMALL PARTICLES

Example:

- (a) Dissolving of salt (or sugar) in water: the water and salt (or sugar) both are made up of particles, the particles of salt get into the spaces between particles of water.
- (b) If we take 2ml of dettol we can detect smell even in very dilute solution.

2. PARTICLES OF MATTER HAVE SPACE BETWEEN THEM

After dissolving 50gm salt in 100 ml of water, volume of solution remains 100 ml.

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just

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