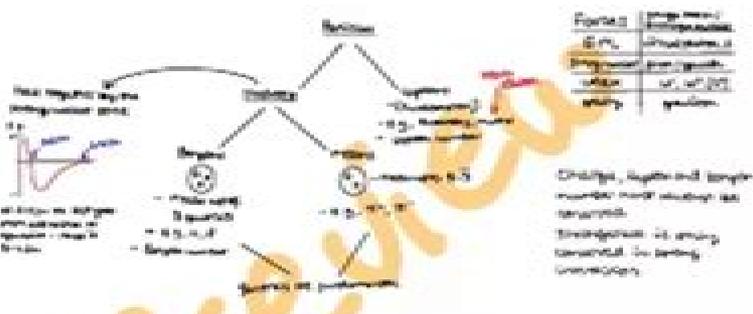


# Particles and Quantum



To absorb must be energy...  $E_{photon} = h\nu$

Annihilation - pair of particles and its corresponding antiparticle created and their rest energy is converted into photons

$$e^- + e^+ \rightarrow \gamma + \gamma$$

Pair production - when photon turns into a particle and its corresponding antiparticle. The photon must have at least the sum amount of energy which rest energies of particles

$$h\nu \geq 2m_0c^2$$

**Excitation**

Exciting means they gain energy through energy to climb their atomic

Alpha particles - Helium nucleus,  ${}^4_2\text{He}$

- Highly ionizing, so it is dangerous
- Relatively penetrating - stopped by a piece of paper or foil on their

$${}^{238}_{92}\text{U} \rightarrow {}^{234}_{90}\text{Th} + {}^4_2\text{He}$$

Beta - fast moving electron  $\rightarrow e^-$

- Medium penetrating and velocity slowing - stopped by a few cm of aluminium
- Beta decay -  ${}^Z_C \rightarrow {}^Z_C + e^- + \bar{\nu}_e$  - **neutrino** is almost undetectable



Gamma - high energy EM ray or photon emitted from a high energy nucleus  $\rightarrow \gamma$

- Highly penetrating - stopped by a few m of concrete

Isotopes - same element - same number of protons

- different relative atomic mass - different number of neutrons

$${}^A_Z X \rightarrow {}^A_Z X + \gamma$$

Specific charge -  $\frac{Q}{m}$  (C/kg)

- charge to mass ratio

Electron volt - charge of an electron when accelerated through a pot of one volt

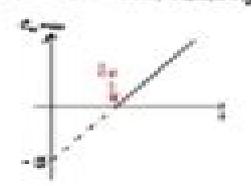
- $1.6 \times 10^{-19} \text{ J}$
- $1 \text{ eV}$



## Photoelectric effect

Use photons of sufficient energy or absorbed by electrons in the surface of a metal, liberating them

- $E_{photon} = hf = \phi + E_{kmax}$
- 1 - photon absorbed -  $hf$  (total energy)
- 2 - frequency
- 3 - work function - minimum energy to liberate electrons
- 4 -  $E_{kmax}$  - kinetic energy of electrons
- 5 - stopped potential - minimum potential needed for electrons to be liberated



The photoelectric effect provides the particle nature of light due to the fact that the continuous function between photons and electrons is not observed experimentally. In fact there is a sharp line indicating the intensity of light cannot liberate electrons if its frequency is below the threshold frequency. This is because the electrons are liberated only if the photon energy is greater than the threshold potential.

Stopping potential - the pot. needed to stop any electrons from reaching the other side

- $E_{kmax} = eV_s$
- 6 - change of potential
- 7 - stopping potential

## De Broglie

The de Broglie wavelength is the wavelength that a particle can have. The wave nature of a particle will occur by being described as a packet of waves which propagate through space and time. This is because the electron is displaying wave-like behavior.



## Fluorescent tubes

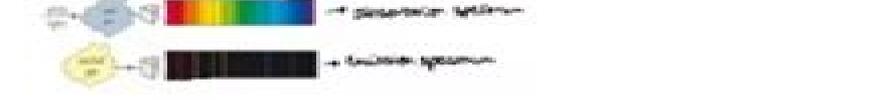


These tubes contain mercury under a high voltage is applied. This excites the gas molecules to which some of the mercury atoms emit photons when they relax. When the gas molecules collide with the tube wall, they are excited to higher energy levels when the electrons relax, they emit photons in the UV range. A phosphor coating in the tube absorbs the photons and emit visible light.

## Energy levels

Electrons can be excited to higher energy levels either by absorbing photons from the light or by the collision. In a discharge tube or in a gas discharge cell when ionizing the electron energy.

Ionization energy - is enough energy to pull an electron out of an atom, so will completely change the atom.



# A Level Physics Notes

**Marcel A. Müller**



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