

# Atomic structure and the periodic table

## Atomic structure



protons  
neutrons  
electrons

atoms have no overall charge

particle	relative mass	relative charge
proton	1	+1
neutron	1	0
electron	very small	-1

## elements

elements - substances made up of atoms with the same atomic number

isotopes of an element - atoms with the same number of protons but different numbers of neutrons

relative atomic mass - the average mass number for an element

## Equations, mixtures and chromatography

### chemical formulas and equations

chemical formula - shows the proportion of atoms of each element in a compound

chemical equation - shows the overall change in a reaction

example:

reactants

products

word equation: methane + oxygen → carbon dioxide + water

symbol equation:  $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$

chromatogram

insoluble components stay on the base line



submerge it

components separate as they move up the paper

## nuclear symbols

mass number = total number of protons and neutrons in an atom



atomic number = number of protons in an atom

## compounds

compound - a substance formed from two or more elements chemically bonded together in fixed proportions

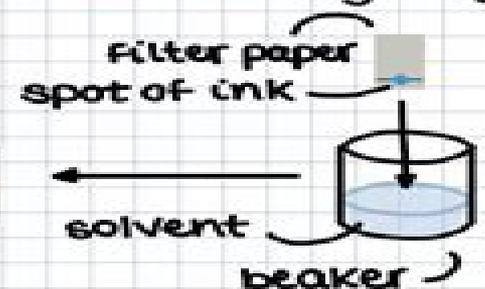
molecule - particle containing two or more non-metal atoms bonded covalently. Molecules can be elements or compounds

## mixtures

mixtures - substances made up of different elements or compounds that are not chemically bonded to each other

mixtures can be separated using physical methods

## paper chromatography



filter paper

spot of ink

solvent

beaker

# Atomic Structure Guided Notes Answers

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## Atomic Structure Guided Notes Answers:

**Edexcel AS/A Level Year 1 Chemistry Student Guide: Topics 1-5** George Facer, Rod Beavon, 2015-10-09 Exam Board Edexcel Level AS A level Subject Chemistry First Teaching September 2015 First Exam June 2016 Reinforce students understanding throughout their course with clear topic summaries and sample questions and answers to help your students target higher grades Written by experienced examiners George Facer and Rod Beavon our Student Guides are divided into two key sections content guidance and sample questions and answers Content guidance will Develop students understanding of key concepts and terminology this guide covers topics 1 5 atomic structure and the periodic table bonding and structure redox 1 inorganic chemistry and the periodic table formulae equations and amounts of substance Consolidate students knowledge with knowledge check questions at the end of each topic and answers in the back of the book Sample questions and answers will Build students understanding of the different question types so they can approach questions from topics 1 5 with confidence Enable students to target top grades with sample answers and commentary explaining exactly why marks have been awarded

OCR A Chemistry AS Student Unit Guide: F321 New Edition Atoms, Bonds and Groups Mike Smith, 2012-08-31 Written by a former senior examiner Mike Smith this OCR A AS Chemistry Student Unit Guide is the essential study companion for Unit F321 Atoms Bonds and Groups This full colour book includes all you need to know to prepare for your unit exam clear guidance on the content of the unit with topic summaries knowledge check questions and a quick reference index examiner s advice throughout so you will know what to expect in the exam and will be able to demonstrate the skills required exam style questions with graded student responses so you can see clearly what is required to get a better grade

Study Guide and Student Solutions Manual to Accompany Physics for Scientists and Engineers, by Serway Raymond A. Serway, Steven Van Wyk, 1996

**A Reader's Guide to Contemporary Literature** William Swan Sonnenschein, 1895

*A Reader's Guide to the Choice of the Best Available Books (about 50,000) in Every Department of Science, Art & Literature, with the Dates of the First & Last Editions, & the Price, Size & Publisher's Name of Each Book* William Swan Sonnenschein, 1901

*Study Guide, Student Solutions Manual* John R. Gordon, 1998

**The Brass World and Platers Guide**, 1922

**Resources in Education**, 1990-07

**The Official Guide for GMAT Review**, 1988

*Platers' Guide*, 1922

**The Chemical News and Journal of Industrial Science**, 1906

**Study Guide/Selected Solutions Manual** Julie R. Frentrup, 2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book

The Praxis Series Official Study Guide: Middle School: Science,

**Isis** George Sarton, 1924 Brief table of contents of vols I XX in v 21 p 502 618

**Instructor's Guide for Introductory Chemistry in the Laboratory** James F. Hall, 1996

*Teacher's Manual and Resource Guide for Exploring the Sciences* Herbert Drapkin, 1964

*The Journal of Education*, 1901

**Reader's Guide to Periodical Literature Supplement**, 1922

Research in Education, 1973

**Readers' Guide to Periodical Literature**

,1920

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