

Temperature, Energy, and Heat Study Guide

I. Basic Definitions

- **Temperature:**
 - A measure of the average kinetic energy of the particles in a substance.
 - Indicates how hot or cold something is.
 - Measured in Celsius ($^{\circ}\text{C}$), Fahrenheit ($^{\circ}\text{F}$), or Kelvin (K).
 - SI unit is Kelvin (K).
- **Heat:**
 - The transfer of thermal energy between objects at different temperatures.
 - Flows from hotter to colder objects.
 - Measured in Joules (J).
- **Thermal Energy:**
 - The total kinetic and potential energy of all the particles in a substance.
 - Depends on temperature and the amount of substance.
- **Absolute Zero:**
 - The theoretical temperature at which all molecular motion stops (0 K, -273.15°C).

II. Temperature Scales

- **Celsius ($^{\circ}\text{C}$):**
 - Water freezes at 0°C and boils at 100°C .
- **Fahrenheit ($^{\circ}\text{F}$):**
 - Water freezes at 32°F and boils at 212°F .
- **Kelvin (K):**
 - Absolute temperature scale.
 - 0 K is absolute zero.[†]
 - $\text{K} = ^{\circ}\text{C} + 273.15$

III. Heat Transfer

- **Conduction:**
 - Transfer of heat through direct contact.
 - Occurs primarily in solids.
 - Good conductors (metals) transfer heat quickly.
 - Insulators (wood, plastic) resist heat transfer.
- **Convection:**
 - Transfer of heat through the movement of fluids (liquids or gases).
 - Hotter, less dense fluids rise; cooler, denser fluids sink.
 - Creates convection currents.

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Rosina Ehmman



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