

## Alberta Exam Chem 30 unit 4 Version 2 ANSWER KEY

1. In the acid-base reaction



- A) reactants are favoured
- B)  $\text{HC}_6\text{H}_5\text{O}_2^-(\text{aq})$  is acting as an acid
- C)  $\text{SO}_4^{2-}(\text{aq})$  is a proton acceptor
- D)  $\text{HSO}_4^-(\text{aq})$  and  $\text{HC}_6\text{H}_5\text{O}_2^-(\text{aq})$  are a conjugate acid-base pair

**Incorrect.** Your answer=A, Correct answer=C

**Explanation:** In the above reaction, the conjugate acid base pair is  $\text{HSO}_4^-$  (acting as an acid, proton donor) and  $\text{SO}_4^{2-}$  (conjugate base, proton acceptor). Choice C is a correct answer.

2.

Tooth decay results from the dissolving of tooth enamel, or hydroxyapatite,  $\text{Ca}_5(\text{PO}_4)_3\text{OH}(\text{s})$ . Bacteria in the presence of sugars in the mouth change the pH of the mouth which causes the following equilibrium:



When sugar ferments on teeth and gums, the mouth becomes more acidic. Tooth enamel

- A) is formed as the concentration of the hydroxide ion decreases and the equilibrium shifts to the left
- B) is formed as the concentration of the hydroxide ion increases and the equilibrium shifts to the left
- C) breaks down as the concentration of the hydroxide ion decreases and the equilibrium shifts to the right
- D) breaks down as the concentration of the hydroxide ion increases and the equilibrium shifts to the right

**Correct.** Your answer=C, Correct answer=C

**Explanation:** Using your common sense and remembering your last visit to the dentist can help you out here.  $\text{Ca}_5(\text{PO}_4)_3\text{OH}(\text{s})$ , tooth enamel, breaks down in an acid environment. The  $\text{H}^+$  ions would neutralize the  $\text{OH}^-$  ions and the decrease would shift the equilibrium to the right, breaking down the enamel. Only Choice C is correct.

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