

Gas Variables

How are the variables that describe a gas related?

Why?

Imagine buying a balloon bouquet at a party store. How will the helium gas in the bouquet behave if you carry it outside on a hot summer day? How will it behave if you carry it outside during a snowstorm? What happens if the balloons are made of latex, which can stretch? What happens if the balloons are made of Mylar[®], which cannot stretch? What if you add just a small amount of gas to each balloon? What if you add a lot of gas? In this activity, you will explore four variables that quantify gases—pressure (P), volume (V), temperature (T), and moles (n) of gas. These four variables can be related mathematically so that predictions about gas behavior can be made.

Model 1 – Gases in a Nonflexible Container

Experiment A (Adding more gas)



A1

Volume = 1 unit
 External pressure = 1 atm
 Internal pressure = 1 atm
 Temperature = 200 K



A2

Volume = 1 unit
 External pressure = 1 atm
 Internal pressure = 2 atm
 Temperature = 200 K



A3

Volume = 1 unit
 External pressure = 1 atm
 Internal pressure = 3 atm
 Temperature = 200 K

Experiment B (Heating the gas)



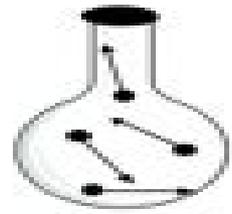
B1

Volume = 1 unit
 External pressure = 1 atm
 Internal pressure = 1 atm
 Temperature = 200 K



B2

Volume = 1 unit
 External pressure = 1 atm
 Internal pressure = 2 atm
 Temperature = 400 K



B3

Volume = 1 unit
 External pressure = 1 atm
 Internal pressure = 3 atm
 Temperature = 600 K

*Note: Volume in this model is recorded in units rather than liters because 4 molecules of gas at the conditions given would occupy a very small space ($\sim 1 \times 10^{-22}$ gL). The particles shown here are much larger compared to the space between them than actual gas particles.



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Robert R. Redfield



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