

## MCQ Module - Understanding Chemical Equilibrium Practice

### Instructions:

Answer the three multiple-choice questions below, choosing the correct response from the four options given. You will receive one mark for each correct answer.

### Question 1/10

1. What does the symbol  $K_c$  represent?

- a) The equilibrium constant for a reaction
- b) The concentration of a reactant in a reaction
- c) The concentration of a product in a reaction
- d) The rate constant for a reaction

2. What is the value of  $K_c$  for the reaction  $2\text{H}_2\text{O}(\text{l}) \rightleftharpoons 2\text{H}_2(\text{g}) + \text{O}_2(\text{g})$  at 25°C?

- a)  $1.0 \times 10^{-13}$
- b)  $1.0 \times 10^{-14}$
- c)  $1.0 \times 10^{-15}$
- d)  $1.0 \times 10^{-16}$

3. What is the value of  $K_c$  for the reaction  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$  at 25°C?

- a)  $1.0 \times 10^{-5}$
- b)  $1.0 \times 10^{-6}$
- c)  $1.0 \times 10^{-7}$
- d)  $1.0 \times 10^{-8}$

4. What is the value of  $K_c$  for the reaction  $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$  at 25°C?

- a)  $1.0 \times 10^{-2}$
- b)  $1.0 \times 10^{-3}$
- c)  $1.0 \times 10^{-4}$
- d)  $1.0 \times 10^{-5}$

5. What is the value of  $K_c$  for the reaction  $\text{H}_2(\text{g}) + \text{Cl}_2(\text{g}) \rightleftharpoons 2\text{HCl}(\text{g})$  at 25°C?

- a)  $1.0 \times 10^{-1}$
- b)  $1.0 \times 10^{-2}$
- c)  $1.0 \times 10^{-3}$
- d)  $1.0 \times 10^{-4}$

# Basic Ict Multiple Choice Questions And Answers

**Camilla Rothe**



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