

Answers to The Oceans End of Unit Test

Q	Answer with marks	Marking suggestions
1(a)	CO ₂ is a gas (1); which has more entropy as there are more ways of arranging its molecules (1)	
1(b)	$\Delta S_{\text{sys}}^{\circ} = 214 + 38 - 93$ (1); $= +159 \text{ J K}^{-1} \text{ mol}^{-1}$ (1), including sign and units	
1(c) (i)	$-180 \times 1000/300$ (1); $= -600 \text{ J K}^{-1} \text{ mol}^{-1}$ (1), including sign and units	
1(c) (ii)	-150 (1) ($\text{J K}^{-1} \text{ mol}^{-1}$)	
1(d)	$\Delta S_{\text{tot}}^{\circ} = \Delta S_{\text{sys}}^{\circ} + \Delta S_{\text{sur}}^{\circ}$ This is negative at 300K and positive at 1200K (1); Positive values means reactions are possible (1)	
1(e) (i)	<p>Top line correct, labelled ΔH_f or enthalpy change of formation (1); Rest of cycle correct, arrows labelled ΔH_f or some attempt to assign specific enthalpy changes (1)</p>	Or alternative cycle drawn as an enthalpy level diagram
1(e) (ii)	Lattice enthalpy $= \Delta H_f - \Delta H_f - \Delta H_f - \Delta H_f$ (1) $= -635 - (178 + 249) - (596 + 1152 - 147 + 753)$ $= -3416 \text{ kJ mol}^{-1}$ (1)	Units not essential for mark
1(f) (i)	At the bottom of the lake the pressure is higher (1); this means the equilibrium will move to the left, increasing the concentration of CO ₂ (aq) (1)	
1(f) (ii)	Deep oceans can absorb a lot of CO ₂ (1); this acts as a 'sink' for the extra CO ₂ we are producing (1)	

Q	Answer with marks	Marking suggestions
2(a)	Ammonia has only (an average) of one hydrogen bond per molecule (1); water has two (1); greater intermolecular forces mean higher boiling point (1)	
2(b)	Ammonia would be less efficient (1); as it has a lower specific heating capacity than water (1); thus less energy would be transferred per mole/gram of ammonia (1)	

Q	Answer with marks	Marking suggestions
3(a) (i)	It is an equilibrium (1); so the benzoic acid is not fully reacted/ ionised (1)	
3(a) (ii)	$K_a = \frac{[\text{C}_6\text{H}_5\text{COO}^-][\text{H}_3\text{O}^+]}{[\text{C}_6\text{H}_5\text{COOH}]}$ (1) for products on nominator; reactants on denominator; (1) completely correct	
3(b) (i)	$-\lg[\text{H}^+]/-\lg[\text{H}_2\text{O}^+]$ (1)	
3(b) (ii)	$[\text{H}_2\text{O}^+]^2 = (1 \times 10^{-14}) \times 0.010 = 1 \times 10^{-16}$ (1) $[\text{H}_2\text{O}^+] = 1 \times 10^{-8} \text{ mol dm}^{-3}$ (1) $\text{pH} = 8$ (1)	
3(b) (iii)	$\text{pH} = 2$ (1)	
3(b) (iv)	HCl is fully ionised (1); (benzoic acid is partially ionised) Hence the concentration of $[\text{H}_2\text{O}^+]$ is lower (1)	

Answers To The Atmosphere End Of Unit Test Benjamin Mills

R Barnett



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