

- 1  $\text{SO}_2$  molecule contains 2 bond pairs (bp) and 1 lone pair (lp) around the S atom. Hence the 3 electron pairs are arranged in a trigonal planar shape at  $120^\circ$  apart. Since lp–bp repulsion is greater than bp–bp repulsion, the bond angle is slightly reduced from  $120^\circ$ . C

2

element	C	H	F
% composition	23.5	2.0	74.5
amount	$\frac{23.5}{12.0} \approx 2$	$\frac{2.0}{1.0} = 2$	$\frac{74.5}{19.0} \approx 4$
mole ratio	1	1	2

D

Molecular formula of HFC–134a must satisfy the empirical formula of  $\text{CHF}_2$ . E.g.  $\text{C}_2\text{H}_2\text{F}_4$ .

- 3
- |   |  |
|---|--|
| <p>A <math>{}_5\text{B}</math>: <math>[\text{He}] 2s^2 2p^3</math> (1 unpaired e)</p> <p>B <math>{}_6\text{C}</math>: <math>[\text{He}] 2s^2 2p^4</math> (2 unpaired e)</p> <p>C <math>{}_{10}\text{Ne}</math>: <math>[\text{Ne}]</math> (0 unpaired e)</p> <p>D <math>{}_7\text{N}</math>: <math>[\text{He}] 2s^2 2p^3</math> (3 unpaired e)</p> | <p><math>{}_{14}\text{Si}</math>: <math>[\text{Ne}] 3s^2 3p^4</math> (2 unpaired e)</p> <p><math>{}_{17}\text{Cl}</math>: <math>[\text{Ne}] 3s^2 3p^5</math> (1 unpaired e)</p> <p><math>{}_{19}\text{K}</math>: <math>[\text{Ar}] 4s^1</math> (1 unpaired e)</p> <p><math>{}_{16}\text{S}</math>: <math>[\text{Ne}] 3s^2 3p^4</math> (2 unpaired e)</p> |
|---|--|
- D

Option B can be eliminated straight away based on the atomic number [C has 11 (not 9) fewer protons than Cl].

- 4
1. van der Waals (branched, compact)
  2. van der Waals (elongated)
  3. intermolecular hydrogen bonding
  4. permanent dipole–permanent dipole attraction (similar e cloud size with polar C–Cl)
- b.p. order :  $3 > 4 > 2 > 1$
- C

- 5
- |                                 |  |
|---------------------------------|--|
| A $2(5) + 6 = 16$ electrons     | $:\text{N} \equiv \text{N} \rightarrow \ddot{\text{O}}:$   |
| B $2(6) = 18$ electrons         | $:\ddot{\text{O}} = \ddot{\text{O}} \rightarrow \ddot{\text{O}}:$  |
| C $3(5) + 1 = 16$ electrons     | $[\text{:N} \equiv \text{N} \rightarrow \ddot{\text{N}}:]^- \longleftrightarrow [\text{:}\ddot{\text{N}} = \text{N} = \ddot{\text{N}}:]^-$       |
| D $2(5) + 4 + 2 = 16$ electrons | $[\text{:N} \equiv \text{C} \rightarrow \ddot{\text{N}}:]^{2-} \longleftrightarrow [\text{:}\ddot{\text{N}} = \text{C} = \ddot{\text{N}}:]^{2-}$ |
- B

- 6
- |   |                                |
|---|--------------------------------|
| $\frac{1}{2}\text{N}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \longrightarrow \text{NO}(\text{g})$ | $\Delta H_f^\circ(\text{NO})$  |
| $\text{NO}(\text{g}) \longrightarrow \frac{1}{2}\text{N}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g})$ | $-\Delta H_f^\circ(\text{NO})$ |
| $2\text{NO}(\text{g}) \longrightarrow \text{N}_2(\text{g}) + \text{O}_2(\text{g})$                      |                                |
- A

# Answers H1 Chemistry Paper

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