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PART - C

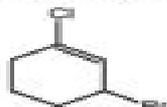
96. HBr reacts with  $\text{CH}_2 = \text{CH} - \text{OCH}_3$  under anhydrous conditions at room temperature to give  
 (1)  $\text{CH}_3\text{CHO}$  and  $\text{CH}_3\text{Br}$  (2)  $\text{BrCH}_2\text{CHO}$  and  $\text{CH}_3\text{OH}$   
 (3)  $\text{BrCH}_2 - \text{CH}_2 - \text{OCH}_3$  (4)  $\text{H}_2\text{C} - \text{CHBr} - \text{OCH}_3$

Ans. (4)

Sol. Electrophilic addition reaction more favourable.



97. The IUPAC name of the compound shown below is



- (1) 2-bromo-6-chlorocyclohex-1-ene (2) 6-bromo-2-chlorocyclohexene  
 (3) 3-bromo-1-chlorocyclohexene (4) 1-bromo-3-chlorocyclohexene

Ans. (3)

98. The increasing order of the rate of HCN addition to compounds A – D is

- (A)  $\text{HCHO}$  (B)  $\text{CH}_3\text{COCH}_3$   
 (C)  $\text{PhCOCH}_3$  (D)  $\text{PhCOPh}$   
 (1)  $\text{A} < \text{B} < \text{C} < \text{D}$  (2)  $\text{D} < \text{B} < \text{C} < \text{A}$   
 (3)  $\text{D} < \text{C} < \text{B} < \text{A}$  (4)  $\text{C} < \text{D} < \text{B} < \text{A}$

Ans. (3)

99. How many moles of magnesium phosphate,  $\text{Mg}_3(\text{PO}_4)_2$ , will contain 0.25 mole of oxygen atoms?

- (1) 0.02 (2)  $3.125 \times 10^{-2}$   
 (3)  $1.25 \times 10^{-2}$  (4)  $2.5 \times 10^{-2}$

Ans. (2)

Sol.  $\text{Mg}_3(\text{PO}_4)_2$

'n' moles  
 $8n = 0.25$

$$n = \frac{0.25}{8} \\ = \frac{25}{8 \times 100} = 3.125 \times 10^{-2}$$

100. According to Bohr's theory, the angular momentum of an electron in 5<sup>th</sup> orbit is

- (1)  $25 \frac{h}{\pi}$  (2)  $1.0 \frac{h}{\pi}$   
 (3)  $10 \frac{h}{\pi}$  (4)  $2.5 \frac{h}{\pi}$

Ans. (4)

Sol.  $mvr = \frac{nh}{2\pi}$   
 $= \frac{5h}{2\pi} = 2.5 \frac{h}{\pi}$

# Anna University Engineering Chemistry 1st Year Notes

**Michael Hoelscher**



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