

Chemistry Student Edition - Basic Answer Key

Chapter 11: Chemical Reactions

Chemical Equations

Questions

1. Early Chinese alchemists discovered an early form of gunpowder. What was the composition of this substance?
2. What later developments were made to the gunpowder recipe that improved its pyrotechnic properties?
3. Make an argument for why the burning of a candle is consistent with the law of conservation of matter/mass.
4. Think of an experiment that you could conduct to demonstrate that mass is conserved for a given chemical change.
5. Balance the following chemical equations:
 - a. $C + O_2 \rightarrow CO$
 - b. $CO + O_2 \rightarrow CO_2$
 - c. $H_2 + Br_2 \rightarrow HBr$
 - d. $K + H_2O \rightarrow KOH + H_2$
 - e. $O_3 \rightarrow O_2$
 - f. $N_2 + H_2 \rightarrow NH_3$
 - g. $Zn + AgCl \rightarrow ZnCl_2 + Ag$
 - h. $Cl_2 + NaI \rightarrow NaCl + I_2$
 - i. $P_4O_{10} + H_2O \rightarrow H_3PO_4$
 - j. $Be_2C + H_2O \rightarrow Be(OH)_2 + CH_4$
 - k. $S + HNO_3 \rightarrow H_2SO_4 + NO_2 + H_2O$
 - l. $NH_3 + CuO \rightarrow Cu + N_2 + H_2O$
 - m. $HCl + CaCO_3 \rightarrow CaCl_2 + H_2O + CO_2$

Answers

1. Early gunpowder was comprised of saltpeter (impure potassium nitrate), sulfur, and charcoal (carbon).
2.
 - a. Purification of potassium nitrate
 - b. Using potassium chlorate instead of potassium nitrate
 - c. Further refinement of the recipe to find the proportions of reactants that yielded the most energetic reaction
3. The burning of a candle obeys the law of conservation of mass because, even though the mass of the candle will decrease as it burns, the mass is not lost. It is transferred into carbon dioxide and water that enter the surrounding air.
4. The reaction between iron and oxygen to produce rust could be studied as an example of conservation of mass. Steel wool could be placed into a covered beaker containing water. The reaction would proceed over time and rust would begin to form. The beaker could be weighed before and after the change. The mass would be the same, even though a new

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Table of Contents Ahc Chapter 11 Chemical Reactions

1. Understanding the eBook Ahc Chapter 11 Chemical Reactions
 - The Rise of Digital Reading Ahc Chapter 11 Chemical Reactions
 - Advantages of eBooks Over Traditional Books
2. Identifying Ahc Chapter 11 Chemical Reactions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ahc Chapter 11 Chemical Reactions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ahc Chapter 11 Chemical Reactions
 - Personalized Recommendations

- Ahc Chapter 11 Chemical Reactions User Reviews and Ratings
- Ahc Chapter 11 Chemical Reactions and Bestseller Lists
- 5. Accessing Ahc Chapter 11 Chemical Reactions Free and Paid eBooks
 - Ahc Chapter 11 Chemical Reactions Public Domain eBooks
 - Ahc Chapter 11 Chemical Reactions eBook Subscription Services
 - Ahc Chapter 11 Chemical Reactions Budget-Friendly Options
- 6. Navigating Ahc Chapter 11 Chemical Reactions eBook Formats
 - ePub, PDF, MOBI, and More
 - Ahc Chapter 11 Chemical Reactions Compatibility with Devices
 - Ahc Chapter 11 Chemical Reactions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ahc Chapter 11 Chemical Reactions
 - Highlighting and Note-Taking Ahc Chapter 11 Chemical Reactions
 - Interactive Elements Ahc Chapter 11 Chemical Reactions
- 8. Staying Engaged with Ahc Chapter 11 Chemical Reactions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ahc Chapter 11 Chemical Reactions
- 9. Balancing eBooks and Physical Books Ahc Chapter 11 Chemical Reactions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ahc Chapter 11 Chemical Reactions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ahc Chapter 11 Chemical Reactions
 - Setting Reading Goals Ahc Chapter 11 Chemical Reactions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ahc Chapter 11 Chemical Reactions
 - Fact-Checking eBook Content of Ahc Chapter 11 Chemical Reactions

- Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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