

Chemistry Worksheets Class 10 on Chapter 1 Chemical Reactions and Equations- Set 1

Q-1: Write the balanced chemical equations for each of the following reactions.

- Liquid hydrogen peroxide decomposes to form water and oxygen gas.
- Acetylene gas burns in oxygen to form carbon dioxide and water along with evolution of heat.
- When solid mercury(II) oxide is heated, liquid mercury and oxygen gas are produced.
- Marble dissolves in a hydrochloric acid to give calcium chloride, water and carbon dioxide.
- Copper sulphate on treatment with potassium iodide precipitates cuprous iodide, liberates iodine gas and also forms potassium sulphate.

Q-2: Fill in the blanks

- _____ is the magnetic oxide of iron.
- In reactivity series, manganese lies _____ zinc
- To the reaction between AgNO_3 solution and NaCl solution, the names given to this type of reactions are _____ and _____.
- Proteins decompose to form _____ in our body.
- If copper turnings are added to colourless AgNO_3 solution, it turns _____ after sometime.

Q-3: Which of the following displacement reactions will not take place?

- $\text{Zn(s)} + \text{FeSO}_4(\text{aq}) \rightarrow \text{ZnSO}_4(\text{aq}) + \text{Fe(s)}$
- $\text{Fe(s)} + \text{CuSO}_4(\text{aq}) \rightarrow \text{FeSO}_4(\text{aq}) + \text{Cu(s)}$
- $\text{Cu(s)} + 2\text{AgNO}_3(\text{aq}) \rightarrow \text{Cu(NO}_3)_2(\text{aq}) + 2\text{Ag(s)}$
- $\text{Cu(s)} + \text{FeSO}_4(\text{aq}) \rightarrow \text{CuSO}_4(\text{aq}) + \text{Fe(s)}$

Q-4: Which of the following statements is false?

- Oxidation results in the loss of an electropositive element and the gain of an electronegative element.
- Reduction results in the loss of an electronegative element and the gain of an electropositive element.
- A species that gains electrons is referred to as a reducing agent.
- An oxidising agent is a type of specie that absorbs hydrogen atoms.

Q-5: Why are certain chemicals and medications kept in coloured bottles?

Q-6: Sakshi desired that her home be white washed. She went to the market and bought 20 kg of quick lime, which she dissolved in 60 litres of water. When she added lime to water, she noticed that the water began to boil even though it was not heated. Give an explanation for her observation. Name the product and write the corresponding chemical equation and word equation.

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Table of Contents Pearson Chemistry Chapters 10 Assessment Answers

1. Understanding the eBook Pearson Chemistry Chapters 10 Assessment Answers
 - The Rise of Digital Reading Pearson Chemistry Chapters 10 Assessment Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Chemistry Chapters 10 Assessment Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pearson Chemistry Chapters 10 Assessment Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Chemistry Chapters 10 Assessment Answers
 - Personalized Recommendations
 - Pearson Chemistry Chapters 10 Assessment Answers User Reviews and Ratings
 - Pearson Chemistry Chapters 10 Assessment Answers and Bestseller Lists

5. Accessing Pearson Chemistry Chapters 10 Assessment Answers Free and Paid eBooks
 - Pearson Chemistry Chapters 10 Assessment Answers Public Domain eBooks
 - Pearson Chemistry Chapters 10 Assessment Answers eBook Subscription Services
 - Pearson Chemistry Chapters 10 Assessment Answers Budget-Friendly Options
6. Navigating Pearson Chemistry Chapters 10 Assessment Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Pearson Chemistry Chapters 10 Assessment Answers Compatibility with Devices
 - Pearson Chemistry Chapters 10 Assessment Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pearson Chemistry Chapters 10 Assessment Answers
 - Highlighting and Note-Taking Pearson Chemistry Chapters 10 Assessment Answers
 - Interactive Elements Pearson Chemistry Chapters 10 Assessment Answers
8. Staying Engaged with Pearson Chemistry Chapters 10 Assessment Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pearson Chemistry Chapters 10 Assessment Answers
9. Balancing eBooks and Physical Books Pearson Chemistry Chapters 10 Assessment Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pearson Chemistry Chapters 10 Assessment Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Pearson Chemistry Chapters 10 Assessment Answers
 - Setting Reading Goals Pearson Chemistry Chapters 10 Assessment Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pearson Chemistry Chapters 10 Assessment Answers
 - Fact-Checking eBook Content of Pearson Chemistry Chapters 10 Assessment Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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