



# COVALENT BONDING

## Practice Problems

Do your answers to the following problems.

### SECTION 5.1 MOLECULAR COMPOUNDS

- Classify each of the following as either ionic or molecular.
  - $\text{NaCl}$
  - $\text{H}_2\text{O}$
  - $\text{HCl}$
  - $\text{KNO}_3$
  - $\text{CH}_4$
  - $\text{H}_2\text{SO}_4$
- Write the formula for the following molecular compounds.
  - Carbon disulfide
  - Carbon tetrachloride
  - Carbon dioxide
  - Carbon monoxide
  - Phosphorus pentachloride
  - Diphosphorus pentoxide

- Write the molecular formula for each of the following compounds.
  - Carbon tetrachloride
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### SECTION 5.2 THE NATURE OF COVALENT BONDING

- Draw the Lewis structure for hydrogen fluoride.  $\text{H}-\text{F}$
- Draw the Lewis structure for carbon dioxide.  $\text{O}=\text{C}=\text{O}$
- Draw the Lewis structure for methane.  $\text{C}$  with four  $\text{H}$  atoms bonded to it.

- Draw the Lewis structure for acetylene.  $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw the Lewis structure for ammonia.  $\text{N}$  with three  $\text{H}$  atoms bonded to it.



### SECTION 5.3 BONDING: THE ORBITALS

- Which orbital and sub-orbital are represented by each of the following?
  - $2s$
  - $2p$
  - $3s$
  - $3p$
  - $4s$
  - $4p$
- Which orbital and sub-orbital are represented by each of the following?
  - $2s$
  - $2p$
  - $3s$
  - $3p$
  - $4s$
  - $4p$

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**Didier Musso**



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