

## 11

## CHEMICAL REACTIONS

## Chapter Test B

## A. Matching

Match each term in Column B with the correct description in Column A. Write the letter of the correct term on the line.

Column A	Column B
_____ 1. a reaction in which an element or compound reacts with oxygen, often producing energy in the form of heat or light	a. decomposition reaction
_____ 2. a reaction in which two or more substances react to form a single substance	b. activity series
_____ 3. an equation that indicates only those particles that actually take part in the reaction	c. spectator ions
_____ 4. a substance that speeds up a reaction without being used up	d. balanced equation
_____ 5. ions that are not directly involved in a reaction	e. double-replacement reaction
_____ 6. a reaction in which a single compound is broken down into two or more products	f. catalyst
_____ 7. an equation in which each side has the same number of atoms of each element	g. combustion reaction
_____ 8. a reaction in which atoms of an element replace the atoms of a second element in a compound	h. combination reaction
_____ 9. a list of metals in order of decreasing reactivity	i. net ionic equation
_____ 10. a reaction that involves an exchange of positive ions between two compounds	j. single-replacement reaction

## B. Multiple Choice

Choose the best answer and write its letter on the line.

- \_\_\_\_\_ 11. In the chemical equation  $2\text{H}_2\text{O}_2(\text{aq}) \xrightarrow{\text{MnO}_2} 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$ , the  $\text{MnO}_2$  is a:
- a. reactant.  
b. product.  
c. spectator ion.  
d. catalyst.
- \_\_\_\_\_ 12. When the equation  $\text{Mg}(\text{s}) + \text{HCl}(\text{aq}) \rightarrow \text{MgCl}_2(\text{aq}) + \text{H}_2(\text{g})$  is balanced, what is the coefficient for HCl?
- a. 1  
b. 2  
c. 3  
d. 4

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**John Robertson Dunlap, Henry Clemens  
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