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 Section 2230-1679 Team _____
 Instructor MESA _____

REPORT SHEET | LAB
**Compounds and
 Their Bonds** | **9**

A. Ions: Transfer of Electrons

Element	Atomic Number	1. Electron Configuration of Atom	2. Electron-Dot Symbol	3. Loss or Gain of Electrons	4. Electron Configuration of Ion	5. Ionic Charge	6. Symbol of Ion	7. Name of Ion
Sodium	11	$1s^2 2s^2 2p^6 3s^1$	Na [•]	lose 1 e ⁻	$1s^2 2s^2 2p^6$	1+	Na ⁺	Sodium
Nitrogen	7	$1s^2 2s^2 2p^3$	$:\ddot{N}:$	gains 3e ⁻	$1s^2 2s^2 2p^6$	3-	N ³⁻	nitride
Aluminum	13	$1s^2 2s^2 2p^6 3s^2 3p^1$	$\cdot\text{Al}\cdot$	lose 3e ⁻	$1s^2 2s^2 2p^6$	3+	Al ³⁺	aluminum
Chlorine	17	$1s^2 2s^2 2p^6 3s^2 3p^5$	$:\ddot{Cl}\cdot$	gains 1e ⁻	$1s^2 2s^2 2p^6 3s^2 3p^6$	1-	Cl ⁻	chloride
Calcium	20	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$	Ca ^{••}	lose 2e ⁻	$1s^2 2s^2 2p^6 3s^2 3p^6$	2+	Ca ²⁺	calcium
Oxygen	8	$1s^2 2s^2 2p^4$	$:\ddot{O}:$	gains 2e ⁻	$1s^2 2s^2 2p^6$	2-	O ²⁻	oxide

* $1s^2 2s^2 2p^6 3s^2 3p^6$

B. Ionic Compounds and Formulas

1. Physical properties

Compound	Appearance	Density	Melting Point
Sodium chloride, NaCl	white crystalline solid	2.17 g/cm ³	1,474 °F

2. Formulas of ionic compounds

Name	Positive Ion	Negative Ion	Formula
Lithium iodide	Li ⁺	I ⁻	LiI
Aluminum oxide	Al ³⁺	O ²⁻	Al ₂ O ₃
Calcium sulfide	Ca ²⁺	S ²⁻	CaS
Magnesium bromide	Mg ²⁺	Br ⁻	MgBr ₂
Potassium nitride	K ⁺	N ³⁻	K ₃ N
Sodium fluoride	Na ⁺	F ⁻	NaF

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Table of Contents Compounds And Their Formulas Report Sheet Answers

1. Understanding the eBook Compounds And Their Formulas Report Sheet Answers
 - The Rise of Digital Reading Compounds And Their Formulas Report Sheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Compounds And Their Formulas Report Sheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Compounds And Their Formulas Report Sheet Answers
 - Personalized Recommendations
 - Compounds And Their Formulas Report Sheet Answers User Reviews and Ratings
 - Compounds And Their Formulas Report Sheet Answers and Bestseller Lists

5. Accessing Compounds And Their Formulas Report Sheet Answers Free and Paid eBooks
 - Compounds And Their Formulas Report Sheet Answers Public Domain eBooks
 - Compounds And Their Formulas Report Sheet Answers eBook Subscription Services
 - Compounds And Their Formulas Report Sheet Answers Budget-Friendly Options
6. Navigating Compounds And Their Formulas Report Sheet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Compounds And Their Formulas Report Sheet Answers Compatibility with Devices
 - Compounds And Their Formulas Report Sheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Compounds And Their Formulas Report Sheet Answers
 - Highlighting and Note-Taking Compounds And Their Formulas Report Sheet Answers
 - Interactive Elements Compounds And Their Formulas Report Sheet Answers
8. Staying Engaged with Compounds And Their Formulas Report Sheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Compounds And Their Formulas Report Sheet Answers
9. Balancing eBooks and Physical Books Compounds And Their Formulas Report Sheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Compounds And Their Formulas Report Sheet Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Compounds And Their Formulas Report Sheet Answers
 - Setting Reading Goals Compounds And Their Formulas Report Sheet Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Compounds And Their Formulas Report Sheet Answers
 - Fact-Checking eBook Content of Compounds And Their Formulas Report Sheet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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