

## Chapter 1 - Introduction

### Classifications of Matter – 1.3

- \* **Matter:** anything that occupies space and has mass
- \* **Chemistry:** the study of matter and its change
- \* **Atom:** smallest distinctive unit
- \* **Molecule:** two or more atoms together
- \* **Substance:** group of identical atoms or molecules
  - o **Element:** atoms are identical
  - o **Compound:** contains different atoms
- \* **Homogenous Mixture:** uniform, seems as one thing
- \* **Heterogeneous Mixture:** non-uniform, able to see the elements



### Physical and Chemical Properties of Matter – 1.4

- \* **Physical Property:** measurable, describes system state
  - o Melting point, solubility, density, color
  - o "Air is a gas," "boiling point of water is 100 degrees," "Sulfur crystals are yellow"
- \* **Chemical Property:** chemical reaction, compound cannot be restored without a change
  - o Burning, explosions, decomposition, rusting
  - o i.e. "Burning Hydrogen in Oxygen to form water," "milk turning sour"



### Measurement – 1.5

- \* **SI base units of Measurement:**
  - o Length – meter (m)
  - o Mass – kilogram (kg)
  - o Time – second (s)
  - o Temperature – Kelvin (K)
  - o Amount of a substance – mole (mol)
- \* **Temperature Scales:**
  - o Kelvin: absolute temperature scale,  $273^{\circ} \text{C} = 1 \text{K}$

### Handling Number – 1.6

- \* **Accuracy:** Correctness – How close the results were to where you intended them to be
- \* **Precision:** Exactness – How close the results were in relation to other results / trials
- \* **Scientific Notation:** Expresses a number as a product of a number between 1 and 10 and the appropriate power of 10
  - o All non-zeros are **significant** (1-9)
  - o A zero between non-zeros is **significant** (209)
  - o Final zeros to the right of the decimal are **significant** (10.0)
  - o Initial zeros are **not significant** (0.47)
  - o Final zeros in a number with no decimal **may or may not be significant** (9,500)
- \* **Adding and subtracting:** same number of digits past the decimal point as the number used with the least precision – the least amount of significant figures
  - o If greater than 5, round up
- \* **Multiplication and Division:** least number of significant figures
- \* **Exact Numbers:** have no effect on significant figures – i.e. 60 seconds in a minute
- \* **Volume (V):** Length cubed,  $\text{mL} = 1 \text{cm}^3$

# Chemistry 101 Study Guide

**Didier Musso**



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