

## CHAPTER 8

### *BONDING: GENERAL CONCEPTS*

#### Questions

15. a. This diagram represents a polar covalent bond as in HCl. In a polar covalent bond, there is an electron rich region (indicated by the red color) and an electron poor region (indicated by the blue color). In HCl, the more electronegative Cl atom (on the red side of the diagram) has a slightly greater ability to attract the bonding electrons than does H (on the blue side of the diagram), which in turn produces a dipole moment.
- b. This diagram represents an ionic bond as in NaCl. Here, the electronegativity differences between the Na and Cl are so great that the valence electron of sodium is transferred to the chlorine atom. This results in the formation of a cation, an anion, and an ionic bond.
- c. This diagram represents a pure covalent bond as in  $H_2$ . Both atoms attract the bonding

# Chapter 8 Bonding General Concepts

**S. Sarg**



## Chapter 8 Bonding General Concepts:

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