

Equilibrium: Assessment Review

Name _____

Skill 1 - The student will be able to draw and explain how a catalyst changes a reaction rate graph and discuss other changes that disrupt an equilibrium system.

1. Using the reaction diagram to the right, representing the reaction: $A + B \rightarrow C + D$ draw a new curve that shows how a catalyst would affect the reaction pathway.

2. Place a checkmark next to the actions that would most likely increase the reaction rate.

- Lowering the temp of the system
- Increasing the concentration of substances present
- Grinding a solid in to fine particles
- Removing a enzyme catalyst



3. Explain why all reactions have activation energy, using your knowledge of collision model.

4. Explain why different reactions have different activation energies.

5. Describe how the activation energy of a reaction affects the overall rate of the chemical reaction differently than a change in concentration, pressure, or temp.

6. A rule of thumb used by most chemists is that the rate of a chemical reaction can be doubled by increasing the reaction temperature by ten degrees Celsius. Explain this drastic increase in reaction rate using your knowledge of the collision model.

Skill 2 - The student will be able to write the equilibrium expression, given a reaction.

When a reaction reaches equilibrium the reverse reaction forms reactants at the same rate that the forward reaction produces products. A reaction may appear to stop when it reaches equilibrium because the amount of products and reactants does not change. This process is called dynamic equilibrium.

A reaction that has reached equilibrium can be described with a ratio known as the equilibrium expression. This expression relates the concentration of products divided by the concentration of reactants, solving for the equilibrium constant, K_c . When writing equilibrium expressions, only aqueous and gaseous substances are included.

Balance the following equations & write the equilibrium expressions needed to solve for K_c :



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