

Name Key

1. Hydrogen,  $H_2$ , is used as rocket fuel. The hydrogen is burned in oxygen,  $O_2$ , according to the following equation. What is the enthalpy change if 200.0 g of hydrogen,  $H_2$  is burned?



$$\frac{200.0 \text{ g } H_2}{2.016 \text{ g } H_2} \times \frac{1 \text{ mol } H_2}{1} = 99.2 \text{ mol } H_2$$

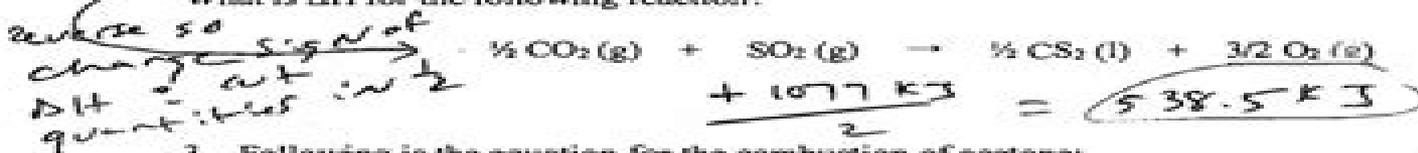
$$99.2 \text{ mol } H_2 \times \frac{-484 \text{ kJ}}{2 \text{ mol } H_2} = -24,006 \text{ kJ}$$

-24,006 kJ  
~~24,006 kJ~~

2. Carbon disulfide burns in air, producing carbon dioxide and sulfur dioxide.



What is  $\Delta H$  for the following reaction?



3. Following is the equation for the combustion of acetone:



How much heat will be given off (in kcal) if 4.5 mol of acetone is burned completely?

$$\frac{-853.6 \text{ kcal}}{2 \text{ mol } C_3H_6O} \times 4.5 \text{ mol } C_3H_6O = -1920.6 \text{ kcal}$$

Exothermic

4. The specific heat of iron is  $0.110 \text{ cal} / \text{g} \cdot ^\circ\text{C}$ . How much heat must be added to 5.2 g of iron at  $25^\circ\text{C}$  to raise its temperature to  $225^\circ\text{C}$ ?

$$\text{heat} = q = m c \Delta T = (5.2 \text{ g}) \left( \frac{0.110 \text{ cal}}{\text{g} \cdot ^\circ\text{C}} \right) (225 - 25^\circ\text{C}) = 114.4 \text{ cal}$$

Endothermic

5. Hydrazine,  $N_2H_4$ , is a colorless liquid used as a rocket fuel. What is the enthalpy change for the process in which hydrazine is formed from its elements (reaction (1) below)? Use Hess's Law to calculate  $\Delta H$  for reaction (1) given the data below.



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**Robert R. Redfield**



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