

Chapter 2 - Practice Test #1

7. An equilibrium system shifts left when the
- A. rate of the forward reaction is equal to the rate of the reverse reaction.
 - B. rate of the forward reaction is less than the rate of the reverse reaction.
 - C. rate of the forward reaction is greater than the rate of the reverse reaction.
 - D. rate of the forward reaction and the reverse reaction are constant.
8. A 1.00 L flask contains a gaseous equilibrium system. The addition of reactants to this flask results in a
- A. shift left and a decrease in the concentration of the products.
 - B. shift left and a increase in the concentration of the products.
 - C. shift right and a decrease in the concentration of the products.
 - D. shift right and a increase in the concentration of the products.
9. Consider the following equilibrium: $\text{CH}_4(g) + \text{H}_2\text{O}(g) + \text{heat} = \text{CO}(g) + 3\text{H}_2(g)$
In which of the following will both stresses shift the equilibrium to the right?
- A. a decrease in temperature and a decrease in volume
 - B. a increase in temperature and a decrease in volume
 - C. a decrease in temperature and a increase in volume
 - D. a increase in temperature and a increase in volume
10. Consider the following equilibrium: $2\text{SO}_2(g) + \text{O}_2(g) = 2\text{SO}_3(g)$ $\Delta H = -198 \text{ kJ}$
There will be no shift in this equilibrium when
- A. more O_2 is added.
 - B. a catalyst is added.
 - C. the volume is increased.
 - D. the temperature is increased.
11. Consider the following equilibrium: $2\text{Fe}_2\text{O}_3 + 3\text{H}_2\text{O}(g) = \text{Fe}_2\text{O}_3 + 3\text{H}_2$
The equilibrium expression is
- A. $\text{Keq} = \frac{[\text{Fe}_2\text{O}_3][\text{H}_2]^3}{[\text{Fe}]^2[\text{H}_2\text{O}]^3}$
 - B. $\text{Keq} = \frac{[\text{Fe}_2\text{O}_3][3\text{H}_2]}{[2\text{Fe}][3\text{H}_2\text{O}]}$
 - C. $\text{Keq} = \frac{[\text{H}_2]^3}{[\text{H}_2\text{O}]^3}$
 - D. $\text{Keq} = [\text{H}_2]^3$
12. Consider the following equilibrium: $\text{N}_2\text{O}_4 = 2\text{NO}_2$ $\text{Keq} = 0.133$
At equilibrium, the $[\text{N}_2\text{O}_4]$ is equal to

Chemical Equilibrium Practice Test Answers

Patrick Vollmar



Chemical Equilibrium Practice Test Answers:

The book delves into Chemical Equilibrium Practice Test Answers. Chemical Equilibrium Practice Test Answers is a vital topic that must be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Chemical Equilibrium Practice Test Answers, encompassing both the fundamentals and more intricate discussions.

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 - Chapter 1: Introduction to Chemical Equilibrium Practice Test Answers
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 - Chapter 3: Chemical Equilibrium Practice Test Answers in Everyday Life
 - Chapter 4: Chemical Equilibrium Practice Test Answers in Specific Contexts
 - Chapter 5: Conclusion
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 5. In chapter 4, the author will scrutinize the relevance of Chemical Equilibrium Practice Test Answers in specific contexts. The fourth chapter will explore how Chemical Equilibrium Practice Test Answers is applied in specialized fields, such as education, business, and technology.
 6. In chapter 5, the author will draw a conclusion about Chemical Equilibrium Practice Test Answers. This chapter will summarize the key points that have been discussed throughout the book.
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