

Chapter 6 Thermochemistry

Homework Problems

- The enthalpy change for the fermentation (without oxygen, O_2 gas) reaction of glucose to alcohol and carbon dioxide is -67 kJ.
 - Write a chemical equation for the fermentation reaction and show how energy is involved in the reaction.
 - Is the reaction exothermic or endothermic?
 - Is energy, in the form of heat, absorbed or evolved as the reaction occurs?
- The specific heat capacity of graphite is 0.71 J/°C·g. Calculate the energy required to:
 - Raise the temperature of 1.0 mol graphite by 1.0 °C.
 - Raise the temperature of 850 g graphite by 150 °C.
 - Raise the temperature of 75 kg graphite from 294 K to 348 K.
- In a coffee-cup calorimeter, 100.0 mL of 1.0 M NaOH and 100.0 mL of 1.0 M HCl are mixed. Both solutions were originally at 24.6 °C. After the reaction, the final temperature is 31.3 °C. Assume that all the solutions have a density of 1.0 g/mL and a specific heat capacity of 4.18 J/g°C. Calculate the enthalpy change for the neutralization of HCl by NaOH. Assume that no heat is lost to the surroundings or to the calorimeter.
- Camphor ($C_{15}H_{26}O$) has an energy of combustion of -5903.6 kJ/mol. A sample of camphor with a mass of 0.1204 g is burned in a bomb calorimeter, the temperature increases by 2.28 °C. Calculate the heat capacity of the calorimeter.
A 0.1964 g sample of quinone ($C_6H_4O_2$) is burned in a bomb calorimeter that has a heat capacity of 1.56 kJ/°C. The temperature of the calorimeter increases by 3.2 °C. Calculate the energy of combustion of quinone per gram and per mole.

Chapter 6 Thermochemistry Homework Problems

Patrick Vollmar



Chapter 6 Thermochemistry Homework Problems:

This book delves into Chapter 6 Thermochemistry Homework Problems. Chapter 6 Thermochemistry Homework Problems is an essential topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Chapter 6 Thermochemistry Homework Problems, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Chapter 6 Thermochemistry Homework Problems
 - Chapter 2: Essential Elements of Chapter 6 Thermochemistry Homework Problems
 - Chapter 3: Chapter 6 Thermochemistry Homework Problems in Everyday Life
 - Chapter 4: Chapter 6 Thermochemistry Homework Problems in Specific Contexts
 - Chapter 5: Conclusion
 2. In chapter 1, the author will provide an overview of Chapter 6 Thermochemistry Homework Problems. The first chapter will explore what Chapter 6 Thermochemistry Homework Problems is, why Chapter 6 Thermochemistry Homework Problems is vital, and how to effectively learn about Chapter 6 Thermochemistry Homework Problems.
 3. In chapter 2, this book will delve into the foundational concepts of Chapter 6 Thermochemistry Homework Problems. The second chapter will elucidate the essential principles that must be understood to grasp Chapter 6 Thermochemistry Homework Problems in its entirety.
 4. In chapter 3, the author will examine the practical applications of Chapter 6 Thermochemistry Homework Problems in daily life. The third chapter will showcase real-world examples of how Chapter 6 Thermochemistry Homework Problems can be effectively utilized in everyday scenarios.
 5. In chapter 4, the author will scrutinize the relevance of Chapter 6 Thermochemistry Homework Problems in specific contexts. The fourth chapter will explore how Chapter 6 Thermochemistry Homework Problems is applied in specialized fields, such as education, business, and technology.
 6. In chapter 5, this book will draw a conclusion about Chapter 6 Thermochemistry Homework Problems. The final chapter will summarize the key points that have been discussed throughout the book.
- This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Chapter 6 Thermochemistry Homework Problems.

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