

CHEMISTRY

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PERIODIC TABLE OF THE ELEMENTS

Diagram showing the periodic table of elements, color-coded by groups. Hydrogen (H) is highlighted with a callout box:

- 1** — atomic number
- H** — atomic symbol
- 1.008** — atomic weight



ATOMIC STRUCTURE

Atomic Number, Z , of an element is the number of protons in the nucleus. For neutral atoms, $Z = +$ the number of electrons. The mass number, A , is the sum of the number of protons and the number of neutrons. Isotopes of an element have the same Z , different A . The atomic mass is the weighted average of isotopes. See also: Atomic Weight. Weighted average of isotopes. See also: Atomic Weight. Weighted average of isotopes. See also: Atomic Weight. Weighted average of isotopes.

ATOMIC QUANTUM NUMBERS AND ORBITALS

Principal quantum number, n , is the number of nodes in the radial wavefunction. It is the number of shells. $n = 1, 2, 3, \dots$. The number of orbitals in a shell is n^2 . The number of orbitals in a subshell is $2l + 1$. The number of orbitals in a subshell is $2l + 1$. The number of orbitals in a subshell is $2l + 1$. The number of orbitals in a subshell is $2l + 1$.

NUCLEAR CHEMISTRY

Nuclear reactions occur when nucleons (protons and neutrons) are rearranged. Nuclear reactions are characterized by the mass number, A , and the atomic number, Z . The mass number, A , is the sum of the number of protons and the number of neutrons. The atomic number, Z , is the number of protons.

MANY-ELECTRON ATOMS

Orbitals and quantum numbers are used to describe the many-electron atoms. The number of orbitals in a shell is n^2 . The number of orbitals in a subshell is $2l + 1$. The number of orbitals in a subshell is $2l + 1$. The number of orbitals in a subshell is $2l + 1$.

n	l	m _l	orbitals
1	0	0	1 orbital
2	0	0	1 orbital
2	1	0, ±1	3 orbitals
3	0	0	1 orbital
3	1	0, ±1	3 orbitals
3	2	0, ±1, ±2	5 orbitals
4	0	0	1 orbital
4	1	0, ±1	3 orbitals
4	2	0, ±1, ±2	5 orbitals
4	3	0, ±1, ±2, ±3	7 orbitals

Atomic configuration: The number of electrons in an atom is equal to the atomic number, Z . The number of electrons in an atom is equal to the atomic number, Z . The number of electrons in an atom is equal to the atomic number, Z .



Chemistry Quickstudy Reference Guides Academic

Lingsheng Yao



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