



# COVALENT BONDING

## Practice Problems

Do your answers to the following problems.

### SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either ionic or molecular.
  - $\text{NaCl}$
  - $\text{H}_2\text{O}$
  - $\text{HCl}$
  - $\text{KNO}_3$
  - $\text{H}_2\text{SO}_4$
  - $\text{H}_2\text{O}_2$
- Write the formula for the following molecular compounds.
  - nitrogen dioxide
  - carbon dioxide
  - carbon monoxide
  - nitrogen monoxide
  - nitrogen trioxide
  - nitrogen pentoxide

- Write the molecular formula for each of the following compounds.
  - nitrogen dioxide
  - carbon dioxide
  - carbon monoxide
  - nitrogen monoxide
  - nitrogen trioxide
  - nitrogen pentoxide
- Write the molecular formula for each of the following compounds.
  - nitrogen dioxide
  - carbon dioxide
  - carbon monoxide
  - nitrogen monoxide
  - nitrogen trioxide
  - nitrogen pentoxide

### SECTION 8.2 THE NATURE OF COVALENT BONDING

- Draw the Lewis structure for hydrogen fluoride.  $\text{H}-\text{F}$
- Draw the Lewis structure for carbon dioxide.  $\text{O}=\text{C}=\text{O}$
- Draw the Lewis structure for methane.  $\text{C}$  with four  $\text{H}$  atoms bonded to it.
- Draw the Lewis structure for ethane.  $\text{H}-\text{C}-\text{C}-\text{H}$  with appropriate  $\text{H}$  atoms.
- Draw the Lewis structure for acetylene.  $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw the Lewis structure for hydrogen cyanide.  $\text{H}-\text{C}\equiv\text{N}$



### SECTION 8.3 BONDING: THE ORBITALS

- Which orbital and sub-orbital are represented by each of the following?
  - $2s$
  - $2p$
  - $3s$
  - $3p$
  - $4s$
  - $4p$
- Which orbital and sub-orbital are represented by each of the following?
  - $2s$
  - $2p$
  - $3s$
  - $3p$
  - $4s$
  - $4p$

# Chapter 8 Covalent Bonding Practice Problems Answers

**DP Hallahan**



## Chapter 8 Covalent Bonding Practice Problems Answers:

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### **Table of Contents Chapter 8 Covalent Bonding Practice Problems Answers**

1. Understanding the eBook Chapter 8 Covalent Bonding Practice Problems Answers
  - The Rise of Digital Reading Chapter 8 Covalent Bonding Practice Problems Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 8 Covalent Bonding Practice Problems Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 8 Covalent Bonding Practice Problems Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 8 Covalent Bonding Practice Problems Answers
  - Personalized Recommendations
  - Chapter 8 Covalent Bonding Practice Problems Answers User Reviews and Ratings

- Chapter 8 Covalent Bonding Practice Problems Answers and Bestseller Lists
- 5. Accessing Chapter 8 Covalent Bonding Practice Problems Answers Free and Paid eBooks
  - Chapter 8 Covalent Bonding Practice Problems Answers Public Domain eBooks
  - Chapter 8 Covalent Bonding Practice Problems Answers eBook Subscription Services
  - Chapter 8 Covalent Bonding Practice Problems Answers Budget-Friendly Options
- 6. Navigating Chapter 8 Covalent Bonding Practice Problems Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 8 Covalent Bonding Practice Problems Answers Compatibility with Devices
  - Chapter 8 Covalent Bonding Practice Problems Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 8 Covalent Bonding Practice Problems Answers
  - Highlighting and Note-Taking Chapter 8 Covalent Bonding Practice Problems Answers
  - Interactive Elements Chapter 8 Covalent Bonding Practice Problems Answers
- 8. Staying Engaged with Chapter 8 Covalent Bonding Practice Problems Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 8 Covalent Bonding Practice Problems Answers
- 9. Balancing eBooks and Physical Books Chapter 8 Covalent Bonding Practice Problems Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 8 Covalent Bonding Practice Problems Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 8 Covalent Bonding Practice Problems Answers
  - Setting Reading Goals Chapter 8 Covalent Bonding Practice Problems Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 8 Covalent Bonding Practice Problems Answers
  - Fact-Checking eBook Content of Chapter 8 Covalent Bonding Practice Problems Answers
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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