

# CHEMISTRY

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## PERIODIC TABLE OF THE ELEMENTS

Diagram showing the periodic table of elements, color-coded by groups. Hydrogen (H) is highlighted with a callout box:

- 1.008** — atomic weight
- H** — atomic symbol
- 1** — atomic number



### ATOMIC STRUCTURE

Atomic Number,  $Z$ , of an element is the number of protons in the nucleus. For neutral atoms,  $Z = +$  the number of electrons. The mass number,  $A$ , is the sum of the number of protons and the number of neutrons. Isotopes of an element have the same  $Z$ , different  $A$ . The atomic mass is the weighted average of isotopes. See also: Atomic Mass, Atomic Weight, Crystal Structure of Solids, Isotopes for a given element.

### ATOMIC QUANTUM NUMBERS AND ORBITALS

**Principal Quantum Number ( $n$ ):**  $n = 1, 2, 3, \dots$   
**Angular Momentum Quantum Number ( $l$ ):**  $l = 0, 1, 2, \dots, n-1$   
**Magnetic Quantum Number ( $m_l$ ):**  $m_l = -l, -(l-1), \dots, 0, \dots, (l-1), l$   
**Spin Quantum Number ( $m_s$ ):**  $m_s = +\frac{1}{2}, -\frac{1}{2}$

### NUCLEAR CHEMISTRY

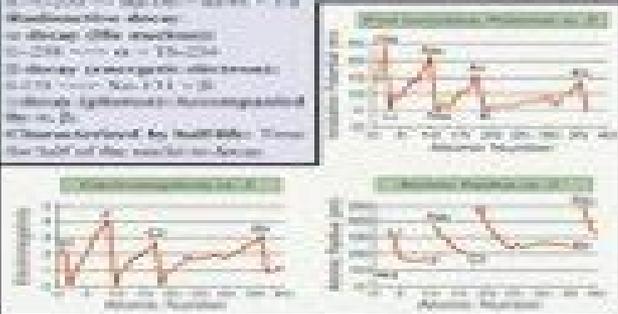
Nuclear reactions alter the number, particles present, or chemical behavior of nuclei and have energy. Energy ( $E$ ) is related to mass ( $m$ ) by the Einstein equation,  $E = mc^2$ .  
**Radioactive Decay:**  $N = N_0 e^{-\lambda t}$   
**Half-life ( $t_{1/2}$ ):**  $t_{1/2} = \frac{\ln 2}{\lambda}$   
**Activity ( $A$ ):**  $A = \lambda N$   
**Decay Constant ( $\lambda$ ):**  $\lambda = \frac{\ln 2}{t_{1/2}}$   
**Decay Energy ( $E$ ):**  $E = \Delta mc^2$   
**Decay Modes:**  $\alpha$ ,  $\beta^-$ ,  $\beta^+$ ,  $\gamma$ ,  $\text{spontaneous fission}$

### MANY-ELECTRON ATOMS

Orbitals and quantum numbers for many-electron atoms are used to describe all atoms, molecules, ions, and solids. A many-electron atom has many subshells with different  $l$  and  $m_l$  values. The main result is that subshells have different energies. The filling of levels is guided by the Aufbau Principle.

$n$	$l$	$m_l$	orbitals
1	0	0	1s
2	0	0	2s
2	1	0	2p <sub>z</sub>
2	1	±1	2p <sub>x</sub> , 2p <sub>y</sub>
3	0	0	3s
3	1	0	3p <sub>z</sub>
3	1	±1	3p <sub>x</sub> , 3p <sub>y</sub>
3	2	0	3d <sub>z<sup>2</sup></sub>
3	2	±1	3d <sub>x<sup>2</sup>-y<sup>2</sup></sub> , 3d <sub>xy</sub>
3	2	±2	3d <sub>xz</sub> , 3d <sub>yz</sub>

**Atomic Configuration:** Shows orbital occupancy. Pauli Exclusion Principle: Each orbital has a unique set of quantum numbers. An orbital may hold up to two electrons, one with spin up,  $m_s = +\frac{1}{2}$ , and one with spin down,  $m_s = -\frac{1}{2}$ .  
**Slater's Rules:** For estimating the energy of an orbital. Shielding constant,  $S$ , is the sum of the shielding constants of all other electrons. See also: Ionization Energy, Atomic Mass, Atomic Number, Atomic Weight, Crystal Structure of Solids, Isotopes for a given element.



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**Yicheng Fang**



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