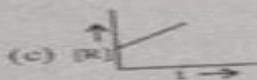
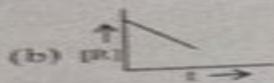
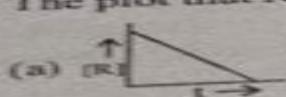


# UNIT 3

## CHEMICAL KINETICS

### Multiple Choice Questions (MCQs)

- The reactions which occurs very fast are
  - mixing aqueous solutions of silver nitrate and sodium chloride
  - Rusting of iron.
  - Inversion of cane sugar
  - hydrolysis of starch.
- In the rate equation, when the conc. of reactants is unity then rate is equal to
  - specific rate constant
  - average rate constant
  - instantaneous rate constant
  - None of above
- The rate of reaction between two specific time intervals is called
  - instantaneous rate
  - average rate
  - specific rate
  - ordinary rate
- Which one of the following statements for the order of a reaction is incorrect ?
  - Order can be determined only experimentally.
  - Order is not influenced by stoichiometric coefficient of the reactants.
  - Order of reaction is sum of power to the concentration terms of reactants to express the rate of reaction.
  - Order of reaction is always whole number.
- Units of rate constant of first and zero order reactions in terms of molarity M unit are respectively
  - $\text{sec}^{-1}, \text{M sec}^{-1}$
  - $\text{sec}^{-1}, \text{M}$
  - $\text{M sec}^{-1}, \text{sec}^{-1}$
  - $\text{M}, \text{sec}^{-1}$
- In a 1<sup>st</sup> order reaction, reactant concentration  $C$  varies with time  $t$  as :
  - $1/C$  increases linearly with  $t$
  - $\log C$  decreases linearly with  $t$
  - $C$  decreases with  $1/t$
  - $\log C$  decreases with  $1/t$
- For a first order reaction, a plot of  $\log(a-x)$  against time is a straight line with a slope equal to
  - $\frac{-k}{2.303}$
  - $-2.303k$
  - $\frac{2.303}{k}$
  - $-\frac{E_a}{2.303R}$
- In a reaction, when the concentration of reactant is increased two times, the increase in rate of reaction was four times. Order of reaction is :
  - zero
  - 1
  - 2
  - 3
- The plot that represents the zero order reaction is :



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**Sebastian Brünink**



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