

# Chapter 1: What is Economics

1. Distribute scarce resources - such as money, land, equipment, or labor
  2. The direct exchange of goods & services without use of money.
  3. Buildings, structures, machines, or tools that are used to produce goods or services.
  4. Items that are used in the production of other goods & services
  5. Finished products that are consumed by an individual.
  6. The ones who buy goods or services for personal use rather than for resale or use in production or manufacturing.
  7. Form of exchange that allows consumers to use items with a promise of repayment over a specified time.
  8. Division of a complex procedure into small tasks, enabling workers to increase output through specialization.
  9. Study of how society chooses to use scarce resources to satisfy its unlimited wants & needs.
  10. Someone who studies economic theory and applies it to the real world.
  11. The production of goods & services using the smallest amounts of resources for the greatest amount of output.
  12. Someone who undertakes and develops a new business enterprise or develops a new product
  13. A resource used to produce goods & services.
  14. Objects or materials that can be purchased to satisfy human wants or needs
  15. Study of an entire economy or one of its principal sectors.
  16. Study of a single factor of an economy - such as individuals, households, businesses, & industries - rather than an economy as a whole.
  17. Any material provided by nature that can be used to produce goods or provide services.
  18. Value lost by rejecting one use of resources in favor of another. In other words, the value of the next-best alternative action that is not taken.
  19. A person, group, or business that makes goods or provides services to satisfy consumers' needs and wants.
  20. Graphic representation showing all of the possible combinations of two goods or services that can be produced in a stated period.
  21. The level of output that results from a given level of input.
  22. The focus of a worker on only one or a few aspects of production in order to improve efficiency.
  23. Scientific & technical techniques used to produce existing products more efficiently or of higher quality.
  24. The sacrifice of one good in order to purchase or produce another
  25. Fundamental condition of economics that results from the combination of limited resources and unlimited wants.
- A. Consumer
  - B. opportunity cost
  - C. Division of Labor
  - D. specialization
  - E. goods
  - F. Entrepreneur
  - G. microeconomics
  - H. scarcity
  - I. Natural Resource
  - J. Credit
  - K. factor of production
  - L. Allocate
  - M. productivity
  - N. Capital Resources
  - O. Capital Goods
  - P. economics
  - Q. producer
  - R. Consumer Goods
  - S. economist
  - T. trade-off
  - U. barter
  - V. Production Possibility Curve
  - W. macroeconomics
  - X. Technology
  - Y. efficiency

# Chapter 9 Economics Test

**Anthony S. Fauci**



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web in the context of a chemical reaction the term neutralization is used for a reaction between an acid and a base or alkali historically this reaction was represented as acid base alkali salt water for example  $\text{hcl naoh nacl h}_2\text{o}$

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web during this reaction the  $\text{h}^+$  ions from the acid and the  $\text{oh}^-$  ions from the alkali react together to form water as this happens energy is released which is detected by the thermometer

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web 1 sodium hydroxide solution is poured into a beaker of hydrochloric acid which contains a thermometer showing room temperature 2 the beaker now contains sodium chloride and water and the

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web reacting two solutions eg acid and alkali place the polystyrene cup inside the glass beaker to make it more stable

measure an appropriate volume of each liquid eg 25 cm<sup>3</sup> place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of sodium hydroxide solution with HCl as 57.9 kJ mol<sup>-1</sup>

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web sep 20 2014 i will illustrate this with some examples when an acid is neutralised by an alkali the reaction is exothermic

eg 1 HCl(aq) + NaOH(aq) → NaCl(aq) + H<sub>2</sub>O(l) for which ΔH = 57 kJ mol<sup>-1</sup> eg 2 HNO<sub>3</sub>(aq) + KOH(aq) → KNO<sub>3</sub>(aq) + H<sub>2</sub>O(l) for which ΔH = 57 kJ mol<sup>-1</sup>

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web the enthalpy change of neutralisation is defined as the enthalpy change when solutions of an acid and alkali react together to produce one mole of water for example HCl(aq) + NaOH(aq) → NaCl(aq) + H<sub>2</sub>O(aq) you can see that one mole of water is produced when one mole of HCl or one mole of NaOH reacts

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