

25. The number of  $\sigma$  bonds in  $N \equiv N$  is
- 1
  - 2
  - 3
  - 4
26. The elements in an ionic compound are held together by
- electrostatic forces of attraction.
  - van der Waals forces
  - the spin of paired electrons.
  - the formation of hybrid orbitals.
  - an electron pair.
27. In every electrolytic and galvanic (voltaic) cell the anode is that electrode
- at which oxidation occurs.
  - which attracts cations.
  - at which electrons are supplied to the solution.
  - at which reduction occurs.
28. Metal X was plated from a solution containing cations of X. The passage of 48.25 C deposited 31 mg of X on the cathode. What is the mass of X (in grams) per mole of electrons?
- 47
  - 62
  - 93
  - 186
29. In a galvanic (voltaic) cell in which the reaction is  $Cd + Cu^{2+} \rightarrow Cu + Cd^{2+}$  and the ions are at unit concentration (activity), the cell potential is
- $$Cd \rightarrow Cd^{2+} + 2e^- \quad 0.4021 \text{ V}$$
- $$Cu \rightarrow Cu^{2+} + 2e^- \quad -0.344 \text{ V}$$
- 0.1383 V
  - 0.4021 V
  - 0.344 V
  - 0.7461 V
  - 0.3677 V
30. In which reaction will an increase in total pressure at constant temperature favor formation of the products?
- $CaCO_3(s) \rightleftharpoons CaO(s) + CO_2(g)$
  - $H_2(g) + Cl_2(g) \rightleftharpoons 2HCl(g)$
  - $2NO(g) + O_2(g) \rightleftharpoons 2NO_2(g)$
  - $COCl_2(g) \rightleftharpoons CO(g) + Cl_2(g)$

Standard Potentials	$E^\circ$
$Mg \rightarrow Mg^{2+} + 2e^-$	2.37V
$Al \rightarrow Al^{3+} + 3e^-$	1.66V
$Zn \rightarrow Zn^{2+} + 2e^-$	0.76V
$Fe \rightarrow Fe^{2+} + 2e^-$	0.44V
$Cu \rightarrow Cu^{2+} + 2e^-$	0.34V
$Ag \rightarrow Ag^+ + e^-$	-0.80V

31. Using only the metals Mg, Al, Zn, Fe, Cu and Ag, together with their 1 M salt solutions, a voltaic cell of the highest possible voltage would be constructed using electrodes of these metals.
- Mg and Ag
  - Mg and Fe
  - Zn and Cu
  - Al and Ag
  - Mg and Al
32.  $E = E^\circ - 0.059/n \log Q$  (Nernst equation)  
 $[H^+] = 1.0 \text{ M}$  initially,  $P_{O_2} = 1.0 \text{ atm}$
- $$4e^- + O_2(g) + 4H^+(aq) \rightleftharpoons 2H_2O(l) \quad E^\circ = 1.23 \text{ V}$$
- Based on the information above, which statement is correct?
- $n = 1$ , since one mole of oxygen is being considered.
  - Addition of base should result in an  $E$  value, which is less than 1.23 V.
  - $E$  is independent of the pH of the solution.
  - $Q = \frac{[H_2O]^2}{[O_2][H^+]^4}$
33. The equilibrium constant for the gaseous reaction  $C + D \rightleftharpoons E + 2F$  is 3.0 at 50 °C. In a 2.0 L flask at 50 °C are placed 1.0 mol of C, 1.0 mol of D, 1.0 mol of E, and 3.0 mol of F. Initially, the reaction will
- proceed at equal rates in both directions.
  - proceed more rapidly to form E and F.
  - proceed more rapidly to form C and D.
  - not occur in either direction.
- | Compound  | $\Delta G_f^\circ$ , kJ/mol |
|-----------|-----------------------------|
| $H_2O(l)$ | -237                        |
| $H_2O(g)$ | -229                        |
34. At 298 K the equilibrium constant for  $H_2(g) + \frac{1}{2} O_2(g) \rightleftharpoons H_2O(l)$
- is larger than the  $K_{eq}$  for  $H_2(g) + \frac{1}{2} O_2(g) \rightleftharpoons H_2O(g)$
  - will have a value of 1.0 at equilibrium.
  - cannot be computed since data on  $O_2$  and

# Acs Standardized Exam General Chemistry Study Guide

**Matthew Bowling**



## **Acs Standardized Exam General Chemistry Study Guide:**

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THAT STUDENTS ENROLLED IN THE COURSE SHOULD LEARN THE BOOK GIVES MANY SAMPLE PROBLEMS WITH STEP WISE SOLUTIONS SO THAT STUDENTS CAN FOLLOW THE MATERIAL EASIER WHEN NECESSARY MATHEMATICAL FORMULAS ARE GIVEN ALL THROUGHOUT TO FACILITATE THE SOLUTIONS TO NUMERICAL PROBLEMS THIS STUDY GUIDE CAN BE USED BY ANY COLLEGE STUDENT ENROLLED IN GEN CHEM I REGARDLESS OF THE TEXT USED ONE OF THE PITFALLS OF MOST TEXTS IS THE EXCESSIVE AMOUNT OF MATERIAL COVERED BUT FAIL TO EMPHASIZE THE MOST IMPORTANT FEATURES OF THE TOPIC COVERED

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