

Name _____

Multiple Choice (10 points each)

Identify the letter of the choice that best completes the statement or answers the question.

- E 1. For a certain process at 27°C, $\Delta G = +210.6$ kJ and $\Delta H = -168.2$ kJ. What is the entropy change for this process at this temperature? Express your answer in the form, $\Delta S =$ _____ J/K.
- a. +628.3 J/K
b. -141.3 J/K
c. +141.3 J/K
d. 1.26×10^3 J/K
e. -1.26×10^3 J/K

Answer:
$$\Delta G = \Delta H - T\Delta S \quad T = 27^\circ\text{C} + 273 = 300\text{K}$$

$$\Delta S = \frac{\Delta H - \Delta G}{T} = \frac{(-168.2\text{kJ}) - (+210.6\text{kJ})}{300\text{K}} = -\frac{1.26\text{kJ}}{\text{K}}$$

$$\left(\frac{-1.26\text{kJ}}{\text{K}}\right)\left(\frac{10^3\text{J}}{\text{kJ}}\right) = \boxed{\frac{-1.26 \times 10^3\text{J}}{\text{K}}}$$

- A 2. Which response includes all of the following substances that have $\Delta H_f^\circ = 0$, and no other substances?

- I. HCl(g)
II. Na(s)
III. HCl(aq)
IV. F₂(g)

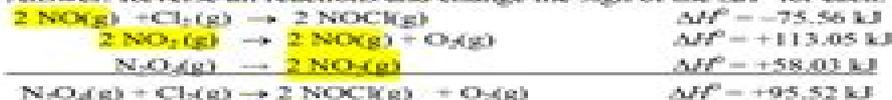
- a. II and IV
b. II
c. I and II
d. I, II, III, and IV
e. I, II, and IV

Answer: All elements in their natural standard state have a $\Delta H_f^\circ = 0$.

- A 3. From the following data at 25°C,

Calculate ΔH° at 25°C for the reaction below.

- a. +95.52 kJ
b. +299 kJ
c. -186.8 kJ
d. -85.52 kJ
e. -246.65 kJ

Answer: Reverse all reactions and change the sign of the ΔH° for each.

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J Spring



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