

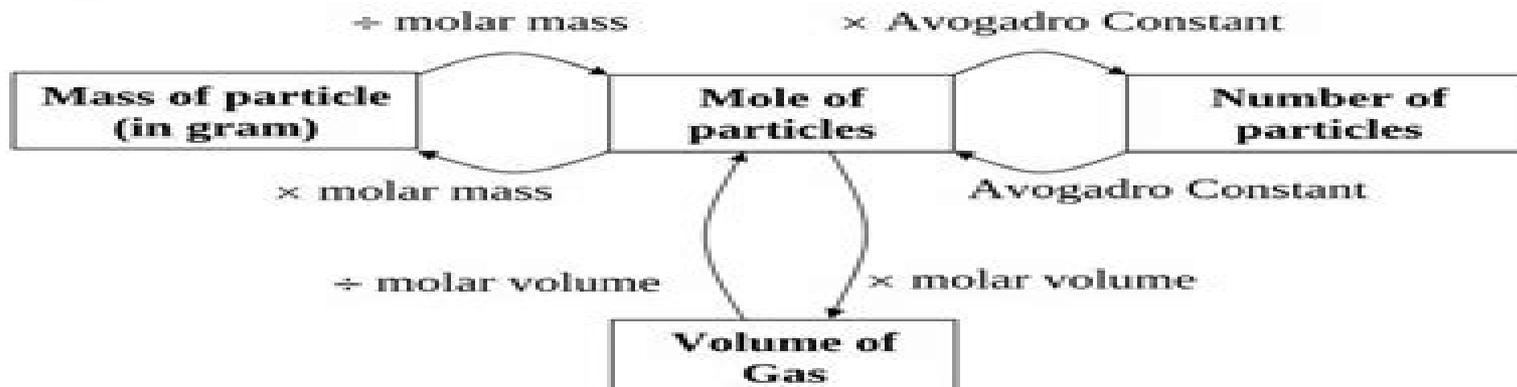
Short Notes: Form 4 Chemistry

Chemical Formulae and Equation

Calculation

<p>For Solid, liquid or gas</p> $\text{number of mole} = \frac{\text{mass of substance}}{\text{molar mass}}$ <p>Molar mass = RAM/RMM/RFM in gram</p>	<p>For gas (only)</p> $\text{number of mole} = \frac{\text{volume of gas}}{\text{molar volume}}$ <p>Molar volume = 24dm³ at room temperature Molar volume = 22.4dm³ at s.t.p.</p>
<p>For Solution</p> $\text{number of mole} = \frac{MV}{1000}$ <p>M = molarity V = Volume of solution in cm³</p>	<p>For quantity of particle(atom,molecule,ion)</p> $\text{number of mole} = \frac{\text{quantity of particle}}{6.02 \times 10^{23}}$

Summary



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marginal revenue marginal cost and profit maximization

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