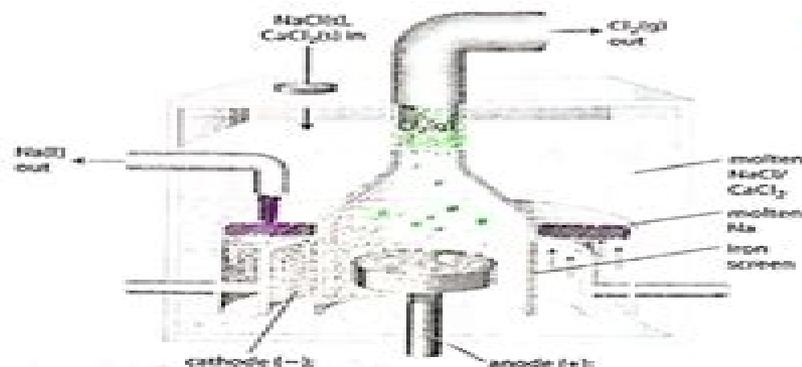


IGCSE – HONORS – ELECTROCHEMISTRY NOTES AND STUDY GUIDE:

Electrolysis of sodium chloride (in the industry)

Industrial electrolysis of molten Sodium Chloride requires extreme temperatures.

- Calcium Chloride is added to the mixture to lower the melting point of Sodium Chloride.
- Molten sodium metal floats in the mixture because of its low density.
- The anode is carbon
- The cathode is iron
- Chlorine gas is collected separately



3.

4. electrolysis of Aluminum oxide (molten)

Electrodes: graphite (carbon)

Electrolyte: molten Alumina (Al₂O₃)

Alumina = Purified Bauxite

Which substance is added to the molten alumina? cryolite

Ions Present: Al³⁺, O²⁻

Half equations:

In the Cathode (Al³⁺ + 3e⁻ → Al) × 4

In the Anode (2O²⁻ → O₂ + 4e⁻) × 3

Overall Equation 4 Al³⁺ + 6 O²⁻ → 4 Al + 3 O₂

Observations:

Cathode: The Molten Aluminum can be observed

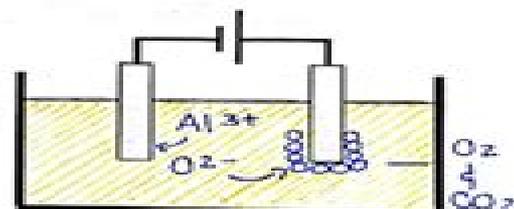
Anode: Bubbles of Oxygen gas will appear

Why is it necessary to change the electrodes periodically?

Because the anodes will burn out. Anode reacts with O₂ to form CO₂

Which is the purpose of the cryolite?

To lower the melting point of the Aluminum Oxide



Chemistry Electrochemistry Study Guide Answers Key

Scott C. Dulebohn



Chemistry Electrochemistry Study Guide Answers Key:

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