

## Chapter 8 Homework

- If  $\Delta H = -54.74 \text{ kJ}$  for a chemical reaction, that reaction \_\_\_\_\_.
  - is exothermic
  - is endothermic
  - cannot occur
  - none of the above
  
- A \_\_\_\_\_ is an experimental set-up that can be used to measure the heat of a chemical reaction or a physical change and determine the specific heat capacity of a substance.
  - manometer
  - thermometer
  - calorimeter
  - whole-body calorimetry
  
- You are asked to design an experiment to measure the enthalpy of a given reaction. The reaction needs to be at atmospheric pressure so which of the following techniques would be ideal for this experiment?
  - bomb manometry
  - bomb calorimetry
  - coffee-cup calorimetry
  - whole-body calorimetry
  
- Given the  $\Delta H^\circ_f$ , calculate the  $\Delta H^\circ_{\text{rxn}}$  for the formation of diamond from graphite.

$$\Delta H^\circ_f (\text{C, graphite}) = 0 \text{ kJ/mol}$$

$$\Delta H^\circ_f (\text{C, diamond}) = 2 \text{ kJ/mol}$$

A. 0 kJ/mol

B. +2 kJ/mol

# Chapter 8 Homework Questions And Problems Answers

1

**Neil Thomson**



## **Chapter 8 Homework Questions And Problems Answers 1:**

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