

## CHAPTER

## STUDY GUIDE

## Covalent Bonding

### Section 8.1 The Covalent Bond

*In your textbook, read about the nature of covalent bonds.*

Use each of the terms below just once to complete the passage.

covalent bond

molecule

sigma bond

exothermic

pi bond

When sharing of electrons occurs, the attachment between atoms that results is called a(n) **(1)** \_\_\_\_\_. When such an attachment is formed, bond dissociation energy is released, and the process is **(2)** \_\_\_\_\_. When two or more atoms bond by means of electron sharing, the resulting particle is called a(n) **(3)** \_\_\_\_\_. If the electrons shared are centered between the two atoms, the attachment is called a(n) **(4)** \_\_\_\_\_. If the sharing involves the overlap of parallel orbitals, the attachment is called a(n) **(5)** \_\_\_\_\_.

*In your textbook, read about single and multiple bonds and bond strength.*

Circle the letter of the choice that best completes the statement or answers the question.

6. In what form do elements such as hydrogen, nitrogen, and oxygen normally occur?
- as single atoms
  - as molecules containing two atoms
  - as molecules containing three atoms
  - as molecules containing four atoms
7. How many electrons are shared in a double covalent bond?
- none
  - one
  - two
  - four
8. Bond length is the distance between
- two molecules of the same substance.
  - the electrons in two attached atoms.
  - the nuclei of two attached atoms.
  - the orbitals of two attached atoms.
9. Which of the following relationships relating to bond length is generally correct?
- the shorter the bond, the stronger the bond
  - the shorter the bond, the weaker the bond
  - the shorter the bond, the fewer the electrons in it
  - the shorter the bond, the lower the bond dissociation energy

# Chapter 8 Covalent Bonding Study Guide Answers

## Pearson

**M Tight**



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concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry

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 solUÇÃO  $\tau = 140 \text{ dias}$   
 $\lambda = \frac{\ln 2}{\tau} = \frac{\ln 2}{140} = 0,0049 \text{ dias}^{-1}$   
 $\ln \frac{A}{A_0} = -\lambda t$   
 $\ln \frac{1}{4} = -0,0049 t$   
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 $\ln 2 = 0,693 \tau$   
 $\tau = \frac{\ln 2}{\lambda}$   
 $\tau = \frac{0,693}{\lambda}$   
 o número n de núcleos de  $^{12}\text{C}$  em 1

g de carbono é 6 02 1023 núcleos mol 12 g mol n 1 g n 6 02 1023 12 5 02 1022 núcleos g o número de núcleos de 14c radioativo é então igual a razão 1 3 10 12

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