

Soluzioni

CAPITOLO 15

VERIFICA LE TUE CONOSCENZE

VALENZA E NUMERO DI OSSIDAZIONE

- 1**
- La valenza di un elemento è il numero di elettroni che l'atomo cede, acquista o condivide quando si lega ad altri atomi.
 - La valenza massima corrisponde al numero romano che contraddistingue il gruppo, quindi, nel caso di un elemento del terzo gruppo della tavola periodica, è 3.
 - Un atomo di carbonio può formare al massimo 4 legami.
- 2** (C) di cariche reali o convenzionali associato a ciascun atomo del composto.
- 3** Regole per il calcolo del numero di ossidazione non corrette [tra parentesi, la versione corretta]:
- l'ossigeno ha sempre n.o. -2 ; [l'ossigeno nei composti ha n.o. -2 , tranne nei perossidi, nei superossidi e nel difluoruro di ossigeno, dove ha rispettivamente n.o. -1 , $-1/2$, $+2$];
 - l'idrogeno ha n.o. $+1$ nei composti e $+2$ nella molecola H_2 ; [l'idrogeno nei composti ha n.o. $+1$ quando è legato a un non metallo e -1 quando è legato a un metallo].

SCRIVERE LE FORMULE DEI COMPOSTI A PARTIRE DAI NUMERI DI OSSIDAZIONE

- 4** Na_2O : in questo modo la somma algebrica dei n.o. è uguale a zero: $2 \cdot (+1) + (-2) = 0$.
- 5** $Ca_3(PO_4)_2$
- 6**
- SO_4^{2-} ; carica dello ione = -2
 - ClO_7^- ; carica dello ione = -1
- 7** La formula Cl_2Ba è sbagliata, in quanto si scrive a sinistra l'atomo con n.o. positivo (Ba_2Cl_2).
- 8** È corretta soltanto: d. SO_2 [tra parentesi, le formule corrette]:
- HS (H_2S)
 - Fe_2S_3 (Fe_2S_3)
 - LiS_2 (Li_2S)

LA NOMENCLATURA DEI COMPOSTI BINARI

- 9** Un composto binario è formato da due elementi diversi. I composti binari si dividono in:
- composti contenenti ossigeno: ossidi (es. ZnO), perossidi (es. H_2O_2), superossidi;
 - composti contenenti idrogeno: idruri (es. NaH), idracidi (es. HF);
 - sali binari (es. $NaCl$, $CaCl_2$).
- 10** In molti composti binari si ha la desinenza *-ure*; tale desinenza non compare negli ossidi.
- 11** La notazione di Stock si utilizza quando gli elementi che formano un composto hanno più di un numero di ossidazione. Comporta l'eliminazione dei prefissi numerici e l'introduzione del numero di ossidazione dell'elemento, come numero romano tra parentesi posto dopo il nome dell'elemento. Si ritiene che in questo modo il nome che contraddistingue l'elemento risulti più breve e più chiaro.
- | | | |
|-------------------|----------|-----------------------|
| (n.o. Fe = $+2$) | $FeBr_2$ | bromuro di ferro(II) |
| (n.o. Fe = $+3$) | $FeBr_3$ | bromuro di ferro(III) |
- 12** Lo ione ossido (O^{2-}) è formato da un atomo di O con n.o. = -2 .
Lo ione perossido (O_2^{2-}) è formato da due atomi di O entrambi con n.o. = -1 , legati da legame covalente.
Lo ione superossido (O_2^-) è formato da due atomi di O con n.o. medio = $-1/2$, legati da legame covalente.
Un ossido acido è formato da un non metallo e ossigeno, un ossido basico è formato da un metallo e ossigeno.
- 13**
- Idrogeno e metallo: idruri (NaH , LiH);
 - idrogeno e non metallo dei gruppi 16 e 17: idracidi (HCl , HF , HBr);
 - metallo e non metallo: sali binari ($NaCl$, $FeCl_2$).
- 14** (B) cianidrico.
- 15** K_2O : ossido di potassio;
 SO_2 : anidride solforosa (n.o. S = $+4$).
 K_2O si chiama *ossido* perché è un ossido basico (metallo e ossigeno).
 SO_2 si chiama *anidride* perché è un ossido acido (non metallo e ossigeno).

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