

Steric # = Total e- pairs on central atom	Electron Distribution	# Bond Pairs	# Lone Pairs	General AXE Formula	Molecular Shape	e.g.	Lewis Structure	3-D Shape (Sketch)	Bond Angle(s)	Polar Molecule? P or NP (4.5)
2	Linear	1	0	AX	Linear	H ₂	H-H		180°	NP
2	Linear	2	0	AX ₂	Linear	BeCl ₂			180°	NP
3	Trigonal Planar	3	0	AX ₃	Trigonal Planar	BCl ₃			120°	NP
		2	1	AX ₂ E	"Bent"	SO ₂			<120°	P or EN >0.4
4	Tetrahedral	4	0	AX ₄	Tetrahedral	CCl ₄			109.5°	NP
		3	1	AX ₃ E	Trigonal Pyramidal	NCl ₃			107°	P or EN >0.4
		2	2	AX ₂ E ₂	Bent	OCl ₂ H ₂ O			104.5°	P or EN >0.4

Dot Vsepr Lab Class Set

Daniel F McAuley



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