

	Reaction 1	Reaction 2	Reaction 3
1. Initial temperature	22.0°C	23.0°C	23.0°C
2. Final temperature	24.0°C	21.0°C	26.0°C
3. The change in temperature, ΔT	+2.0°C	-2.0°C	3.0°C
4. Heat, Q, absorbed by the solution = mass of soln. × ΔT × 1 cal/g °C Note: Mass = sum of masses of solids and liquids used.	151.6g × 2.0°C = 303.2 cal	160 ml × 4.184 cal/g × (-2.0°C) = -640 cal	156.15g - 9.27g = 156.88g × 3.0°C × 1 cal/g °C = 470.64 cal
5. The amount of heat absorbed by the cup (neglect this quantity)	neglect	neglect	neglect
6. TOTAL amount of heat, Q, absorbed	303.2 cal	640 cal	470.64 cal
7. The number of moles of NaOH used in each reaction For solid: gNaOH × 1/MM NaOH For solution: L NaOH × Molarity	0.04 mol	0.04 mol	0.04 mol
8. The heat of reaction per mole of NaOH, ΔH. Include sign. (Line 6/Line 7)	7,580 cal	16,000 cal	11,766 cal

(FOR EACH)

Questions

- Write the net ionic equations for reactions 1, 2, and 3.
- In reaction 1, ΔH₁ represents the heat of solution of NaOH(s). Look at the net ionic equations for reactions 2 and 3 and make a statement concerning the significance of ΔH₂, that is, what has happened to produce the energy change. Also explain the meaning of ΔH₃.
- How does ΔH₂ compare with the sum of ΔH₁ + ΔH₃. Explain.
- Calculate the percent difference between ΔH₂ and the SUM of ΔH₁ + ΔH₃, using ΔH₂ as the reference.

HCl 23.0°C
NaOH 23.0°C

$$\frac{\text{Difference}}{\Delta H_2} \times 100\% =$$

_____ % difference

Answers To Lab 42 Neutralization Reactions

Raffaella Di Napoli



Answers To Lab 42 Neutralization Reactions:

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