

One kilogram of a fluid is compressed according to the relation $(Pv = 0.25)$ where P is in bars and v in m^3/kg . The final volume is $1/4$ the initial volume. Calculate the work done on the fluid and sketch the process on a $P-v$ diagram.

*** solution ***

∴ For closed system

$$W_{1-2} = \int P \cdot dv$$

as this fluid is compressed according to the relation $Pv = 0.25$

but $\rightarrow P$ (bars) & $v \rightarrow \frac{m^3}{kg}$
 $\therefore Pv = 0.25 \times 100 = 25 \frac{kJ}{kg}$

$\therefore P v_0 = 25 \frac{kJ}{kg} \rightarrow P = \frac{25}{v_0} kN$

$$W_{1-2} = \int \frac{25}{v} \cdot dv$$

$$W_{1-2} = 25 \int \frac{dv}{v}$$

$$W_{1-2} = 25 [\ln v]_{v_1}^{v_2}$$

$$= 25 [\ln(v_2) - \ln(v_1)]$$

$$W_{1-2} = 25 \ln\left(\frac{v_2}{v_1}\right)$$

$\therefore v_2 = \frac{1}{4} v_1 \rightarrow \frac{v_2}{v_1} = \frac{1}{4}$
 $\therefore W_{1-2} = 25 \ln\left(\frac{1}{4}\right) = -34.66 \text{ kJ}$
 → which is the specific work done on the system

*** To get work

$$W = m \cdot w_{1-2} = 1 \text{ kg} \times (-34.66) \text{ kJ/kg}$$

$$\therefore W_{1-2} = -34.66 \text{ kJ}$$

(-ve) → work done on the system



2.2 A gas contained within a piston-cylinder assembly expands from an initial volume $V_1 = 0.15 \text{ m}^3$ to a final volume of $V_2 = 0.36 \text{ m}^3$ where the pressure remains constant at 0.5 MPa . Calculate the work, in kJ.

*** solution ***

For closed system shown :-

$$W_{1-2} = \int P \cdot dv$$

→ For constant Pressure Process $P_1 = P_2 = P$

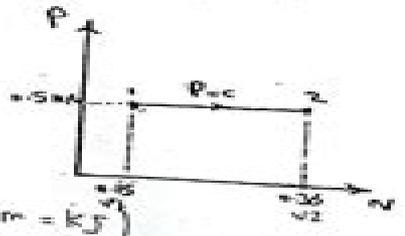
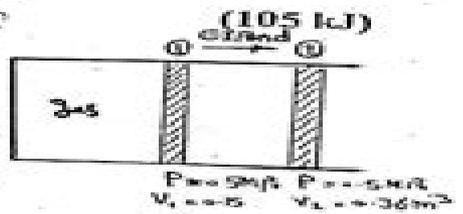
$$\therefore W_{1-2} = P \int dv = P [v]_{v_1}^{v_2}$$

$$= W_{1-2} = P (v_2 - v_1)$$

$$\therefore W_{1-2} = 0.5 \times 10^6 \text{ (N/m}^2) (0.36 - 0.15) \text{ m}^3$$

$$= 105 \text{ kJ} \quad \left(\text{kJ} \cdot \text{m}^3 = \frac{\text{kJ}}{\text{m}^3} \cdot \text{m}^3 = \text{kJ} \right)$$

(+ve) → work done by the system.



Engineering Thermodynamics Solved Problems

James B. Jones, G. A. Hawkins



Engineering Thermodynamics Solved Problems:

200 Solved Problems in Mechanical Engineering Thermodynamics P. E. Liley, 1989 Modern Engineering Thermodynamics - Textbook with Tables Booklet Robert Balmer, 2010-12-20 Modern Engineering Thermodynamics Textbook with Tables Booklet offers a problem solving approach to basic and applied engineering thermodynamics with historical vignettes critical thinking boxes and case studies throughout to help relate abstract concepts to actual engineering applications It also contains applications to modern engineering issues This textbook is designed for use in a standard two semester engineering thermodynamics course sequence with the goal of helping students develop engineering problem solving skills through the use of structured problem solving techniques The first half of the text contains material suitable for a basic Thermodynamics course taken by engineers from all majors The second half of the text is suitable for an Applied Thermodynamics course in mechanical engineering programs The Second Law of Thermodynamics is introduced through a basic entropy concept providing students a more intuitive understanding of this key course topic Property Values are discussed before the First Law of Thermodynamics to ensure students have a firm understanding of property data before using them Over 200 worked examples and more than 1 300 end of chapter problems provide an extensive opportunity to practice solving problems For greater instructor flexibility at exam time thermodynamic tables are provided in a separate accompanying booklet University students in mechanical chemical and general engineering taking a thermodynamics course will find this book extremely helpful Provides the reader with clear presentations of the fundamental principles of basic and applied engineering thermodynamics Helps students develop engineering problem solving skills through the use of structured problem solving techniques Introduces the Second Law of Thermodynamics through a basic entropy concept providing students a more intuitive understanding of this key course topic Covers Property Values before the First Law of Thermodynamics to ensure students have a firm understanding of property data before using them Over 200 worked examples and more than 1 300 end of chapter problems offer students extensive opportunity to practice solving problems Historical Vignettes Critical Thinking boxes and Case Studies throughout the book help relate abstract concepts to actual engineering applications For greater instructor flexibility at exam time thermodynamic tables are provided in a separate accompanying booklet

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