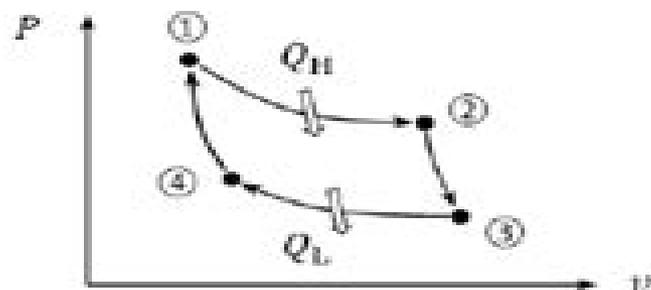


Written Problem

The Carnot cycle is an idealization that involves four processes as shown (see also Figure 6–37 in textbook). As each of the processes is reversible, the entire cycle is also reversible.



- ① → ② : Isothermal expansion at $T = T_H$
- ② → ③ : Adiabatic expansion (no heat transfer)
- ③ → ④ : Isothermal compression at $T = T_L$
- ④ → ① : Adiabatic compression (no heat transfer)

- (a) Suppose the Carnot cycle is carried out in a closed system with an ideal gas as working fluid. Show that

$$Q_H = mRT_H \ln(V_2/V_1)$$

where m is the mass of the ideal gas. (Recall that the internal energy of an ideal gas depends *only* on the temperature T .)

You are now provided with more specifics: The ideal gas is 2 g of air. The Carnot cycle operates between $T_H = 500$ K and $T_L = 300$ K, with $V_2 = 2V_1$.

- (b) What is the amount of heat Q_L (kJ) that is transferred to the cold reservoir in every cycle?
- (c) Determine W_{net} (kJ), the net amount of work that is produced per cycle.

Note that W_{net} can be obtained by summing the $\int P \, dV$ contribution from each of the four processes. This approach, however, is not recommended as there are much easier ways to get to the answer.

Carnot Cycle Numerical Problems With Solutions

Thomas Griffiths



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Table of Contents Carnot Cycle Numerical Problems With Solutions

1. Understanding the eBook Carnot Cycle Numerical Problems With Solutions
 - The Rise of Digital Reading Carnot Cycle Numerical Problems With Solutions
 - Advantages of eBooks Over Traditional Books
2. Identifying Carnot Cycle Numerical Problems With Solutions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Carnot Cycle Numerical Problems With Solutions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Carnot Cycle Numerical Problems With Solutions
 - Personalized Recommendations
 - Carnot Cycle Numerical Problems With Solutions User Reviews and Ratings
 - Carnot Cycle Numerical Problems With Solutions and Bestseller Lists

5. Accessing Carnot Cycle Numerical Problems With Solutions Free and Paid eBooks
 - Carnot Cycle Numerical Problems With Solutions Public Domain eBooks
 - Carnot Cycle Numerical Problems With Solutions eBook Subscription Services
 - Carnot Cycle Numerical Problems With Solutions Budget-Friendly Options
6. Navigating Carnot Cycle Numerical Problems With Solutions eBook Formats
 - ePub, PDF, MOBI, and More
 - Carnot Cycle Numerical Problems With Solutions Compatibility with Devices
 - Carnot Cycle Numerical Problems With Solutions Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Carnot Cycle Numerical Problems With Solutions
 - Highlighting and Note-Taking Carnot Cycle Numerical Problems With Solutions
 - Interactive Elements Carnot Cycle Numerical Problems With Solutions
8. Staying Engaged with Carnot Cycle Numerical Problems With Solutions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Carnot Cycle Numerical Problems With Solutions
9. Balancing eBooks and Physical Books Carnot Cycle Numerical Problems With Solutions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Carnot Cycle Numerical Problems With Solutions
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Carnot Cycle Numerical Problems With Solutions
 - Setting Reading Goals Carnot Cycle Numerical Problems With Solutions
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Carnot Cycle Numerical Problems With Solutions
 - Fact-Checking eBook Content of Carnot Cycle Numerical Problems With Solutions
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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