

CHEMISTRY



Properties of Matter

- Matter is anything that has mass and occupies space.
- Matter is made up of particles.
- Particles are very small.
- Particles are continuously moving.
- Particles attract each other.

States of Matter

- Solid: Particles are closely packed and have a fixed shape and volume.
- Liquid: Particles are less closely packed and have a fixed volume but no fixed shape.
- Gas: Particles are far apart and have no fixed shape or volume.

Changes of State

- Melting: Solid to liquid.
- Freezing: Liquid to solid.
- Boiling: Liquid to gas.
- Condensation: Gas to liquid.
- Sublimation: Solid to gas.
- Deposition: Gas to solid.

Atomic Structure

Atoms are the smallest particles that cannot be created or destroyed in a chemical reaction. They are made up of three sub-particles: electrons, protons, and neutrons.

Electrons

Electrons are negatively charged particles. They are present in all atoms. The number of electrons in an atom is equal to the number of protons.

Protons

Protons are positively charged particles. They are present in all atoms. The number of protons in an atom is equal to the number of electrons.

Neutrons

Neutrons are neutral particles. They are present in all atoms. The number of neutrons in an atom is equal to the number of protons.

Atomic Structure (continued)

Atoms are electrically neutral. The total positive charge of protons is equal to the total negative charge of electrons.

Atomic Models

- Dalton's Atomic Model: Atoms are indivisible particles that cannot be created or destroyed.
- Thomson's Atomic Model: Atoms are made up of positively charged matter in which electrons are embedded.
- Rutherford's Atomic Model: Atoms have a positively charged nucleus in which electrons revolve.

Element	Atomic Number	Atomic Mass
Hydrogen	1	1
Helium	2	4
Lithium	3	7
Boron	5	11
Carbon	6	12
Nitrogen	7	14
Oxygen	8	16
Fluorine	9	19
Neon	10	20

Atomic Structure (continued)

Atoms are made up of three sub-particles: electrons, protons, and neutrons. The number of protons in an atom is equal to the number of electrons.

Atomic Models (continued)

Atoms have a positively charged nucleus in which electrons revolve. The nucleus is made up of protons and neutrons.



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